

HIMACHAL PRADESH BOARD OF SCHOOL EDUCATION, DHARAMSHALA
Model Question Paper
Class – 11th

Duration – 03:00 Hr

Chemistry(Theory)

M.M.: 60

Instructions:

- i) The question paper consists of four sections- A, B, C and D.*
- ii) Internal choices are given in some questions.*
- iii) Section A contains 12 MCQ 1 to 12 of 1 mark each.*
- iv) Section B contains 9 questions from 13 to 21 of 2 marks each.*
- v) Section C contains 6 questions from 22 to 27 of 3 marks each.*
- vi) Section D contains 3 questions from 28 to 30 of 4 marks each.*
- vii) Make neat and clean diagrams where required.*

Section A (1X12)

Q.1 One mole of oxygen molecule is equal to

- a) 6.023×10^{23} atoms of oxygen
- b) 6.023×10^{23} atoms of oxygen
- c) 6.023×10^{23} molecules of oxygen
- d) 6.023×10^{23} molecules of oxygen

Q.2 3d orbitals have values

- a) $n=2, l=3$
- b) $n=3, l=0$
- c) $n=3, l=1$
- d) $n=3, l=2$

Q. 3 Which of the following elements has maximum electron gain enthalpy

- a) F
- b) Cl
- c) Br
- d) I

Q.4 Pyramidal geometry associated with

- a) CH_4
- b) NH_3
- c) H_2O
- d) CO_2

Q.5 Decreasing order of bond angle in H_2O , NH_3 and CH_4

- a) $\text{CH}_4 > \text{NH}_3 > \text{H}_2\text{O}$
- b) $\text{NH}_3 > \text{H}_2\text{O} > \text{CH}_4$
- c) $\text{CH}_4 > \text{H}_2\text{O} > \text{NH}_3$
- d) $\text{H}_2\text{O} > \text{CH}_4 > \text{NH}_3$

Q.6 According to Boyle's law

- a) At constant T, $V \propto P$
- b) At constant T, $V \propto 1/P$
- c) At constant P, $V \propto T$
- d) At constant T, $P \propto T$

Q.7 For the reaction $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$

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