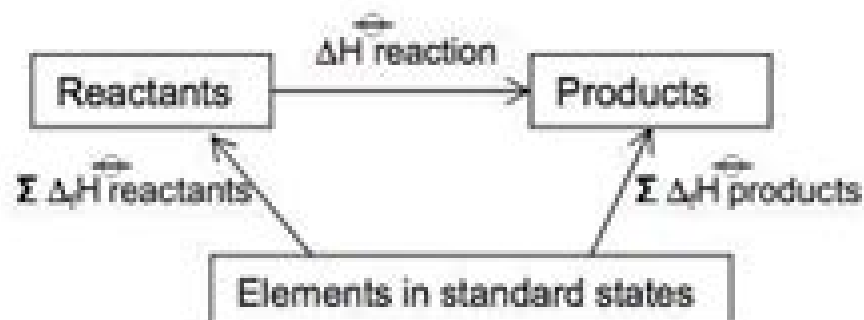


Using Hess's law to determine enthalpy changes from enthalpy changes of formation.

$$\Delta H_{\text{reaction}} = \sum \Delta_f H_{\text{products}} - \sum \Delta_f H_{\text{reactants}}$$



Example 6. What is the enthalpy change for this reaction ?



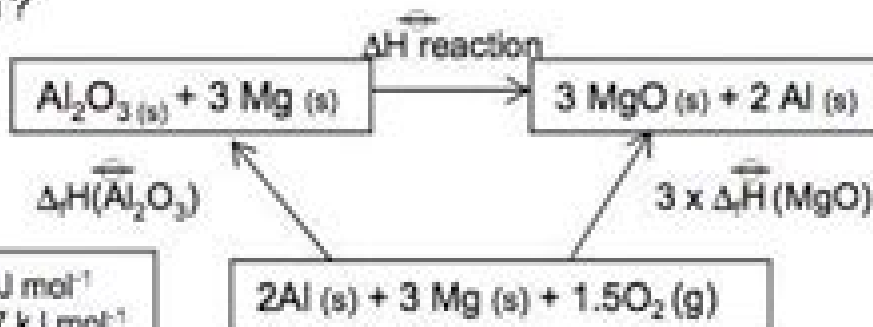
$$\Delta H = \sum \Delta_f H_{\text{products}} - \sum \Delta_f H_{\text{reactants}}$$

$$\Delta H = 3 \times \Delta_f H (\text{MgO}) - \Delta_f H (\text{Al}_2\text{O}_3)$$

$$\begin{aligned} \Delta H &= (3 \times -601.7) - -1675.7 \\ &= -129.4 \text{ kJ mol}^{-1} \end{aligned}$$

Remember
elements have
 $\Delta_f H = 0$

$$\begin{aligned} \Delta_f H (\text{MgO}) &= -601.7 \text{ kJ mol}^{-1} \\ \Delta_f H (\text{Al}_2\text{O}_3) &= -1675.7 \text{ kJ mol}^{-1} \end{aligned}$$



Example 7. Using the following data to calculate the heat of combustion of propene

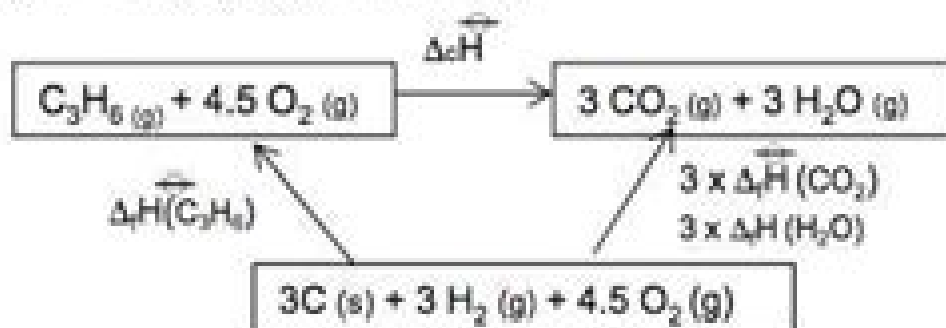
$$\Delta_f H, \text{C}_3\text{H}_6(\text{g}) = +20 \text{ kJ mol}^{-1} \quad \Delta_f H, \text{CO}_2(\text{g}) = -394 \text{ kJ mol}^{-1} \quad \Delta_f H, \text{H}_2\text{O}(\text{g}) = -242 \text{ kJ mol}^{-1}$$



$$\Delta_c H = \sum \Delta_f H_{\text{products}} - \sum \Delta_f H_{\text{reactants}}$$

$$\Delta_c H = [3 \times \Delta_f H (\text{CO}_2) + 3 \times \Delta_f H (\text{H}_2\text{O})] - \Delta_f H (\text{C}_3\text{H}_6)$$

$$\begin{aligned} \Delta_c H &= [(3 \times -394) + (3 \times -242)] - 20 \\ &= -1928 \text{ kJ mol}^{-1} \end{aligned}$$



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Lifelique

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

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1. Understanding the eBook Answers For Chemfax Thermodynamics Enthalpy Lab
 - The Rise of Digital Reading Answers For Chemfax Thermodynamics Enthalpy Lab
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers For Chemfax Thermodynamics Enthalpy Lab
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers For Chemfax Thermodynamics Enthalpy Lab
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers For Chemfax Thermodynamics Enthalpy Lab
 - Personalized Recommendations
 - Answers For Chemfax Thermodynamics Enthalpy Lab User Reviews and Ratings
 - Answers For Chemfax Thermodynamics Enthalpy Lab and Bestseller Lists
5. Accessing Answers For Chemfax Thermodynamics Enthalpy Lab Free and Paid eBooks

- Answers For Chemfax Thermodynamics Enthalpy Lab Public Domain eBooks
- Answers For Chemfax Thermodynamics Enthalpy Lab eBook Subscription Services
- Answers For Chemfax Thermodynamics Enthalpy Lab Budget-Friendly Options
- 6. Navigating Answers For Chemfax Thermodynamics Enthalpy Lab eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers For Chemfax Thermodynamics Enthalpy Lab Compatibility with Devices
 - Answers For Chemfax Thermodynamics Enthalpy Lab Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers For Chemfax Thermodynamics Enthalpy Lab
 - Highlighting and Note-Taking Answers For Chemfax Thermodynamics Enthalpy Lab
 - Interactive Elements Answers For Chemfax Thermodynamics Enthalpy Lab
- 8. Staying Engaged with Answers For Chemfax Thermodynamics Enthalpy Lab
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers For Chemfax Thermodynamics Enthalpy Lab
- 9. Balancing eBooks and Physical Books Answers For Chemfax Thermodynamics Enthalpy Lab
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers For Chemfax Thermodynamics Enthalpy Lab
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers For Chemfax Thermodynamics Enthalpy Lab
 - Setting Reading Goals Answers For Chemfax Thermodynamics Enthalpy Lab
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers For Chemfax Thermodynamics Enthalpy Lab
 - Fact-Checking eBook Content of Answers For Chemfax Thermodynamics Enthalpy Lab
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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