

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive an empirical formula from a chemical structure without performing stoichiometric calculations.
2. To calculate the number of moles of a solid from a weighing measurement in the lab.
3. To determine the molar mass of a solid.
4. To determine the stoichiometric amount of a solid reagent that would be produced.
5. To compare the theoretical amount to the actual amount of substance formed and calculate the percent yield.

Reagents and Apparatus:

Reagent of Sodium Chloride solution: 0.10 mol

Reagent of Calcium Chloride solution: 0.10 mol

Reagent of Silver Nitrate: 0.10 g

Reagent of Silver Nitrate: 0.10 g

Reagent of Silver Nitrate: 0.10 g

Reagent of Silver Nitrate: 0.10 g

Procedure:

Reagents and Calculations:

1. Calculations:

$$\text{moles of sodium chloride} = 0.10 \text{ mol} \times \frac{1.0}{100} = 0.0010 \text{ mol}$$

$$\text{moles of calcium chloride} = 0.10 \text{ mol} \times \frac{1.0}{100} = 0.0010 \text{ mol}$$

$$\text{moles of silver nitrate produced} = 0.0010 \text{ mol} \times \frac{1.0}{100} = 0.0010 \text{ mol}$$

$$\text{moles of silver chloride produced} = 0.0010 \text{ mol} \times \frac{1.0}{100} = 0.0010 \text{ mol}$$

2. Chemical Equations:

3. Reagents:

Reagent of Silver Nitrate: 0.10 g

Reagent of Silver Nitrate: 0.10 g

$$\text{moles of silver chloride produced} = 0.0010 \text{ mol} \times \frac{1.0}{100} = 0.0010 \text{ mol}$$

Answers For Lab 6b Heath Chemistry

Marcel A. Müller



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