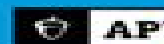


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AP Chemistry

Session 5 – MCQ (Set-Based Questions)



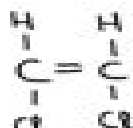
Questions 1-3 refer to the reactions represented below, which are involved in a demonstration commonly known as “underwater fireworks.”



- Ca(OH)₂ precipitates when a 1.0 g sample of CaC₂(s) is added to 1.0 L of distilled water at room temperature. If a 0.064 g sample of CaC₂(s) (molar mass 64g/mol) is used instead and all of it reacts, which of the following will occur and why? (The value of K_{sp} for is 8.0×10^8)
 - will precipitate because $Q > K_{sp}$
 - will precipitate because $Q < K_{sp}$
 - will not precipitate because $Q > K_{sp}$
 - ☒ will not precipitate because $Q < K_{sp}$
- Reaction 2 occurs when an excess of 6 M HCl(aq) solution is added to 100. mL of NaOCl(aq) of unknown concentration. If the reaction goes to completion and 0.010 mol of Cl₂(g) is produced, then what was the molarity of the NaOCl(aq) solution?

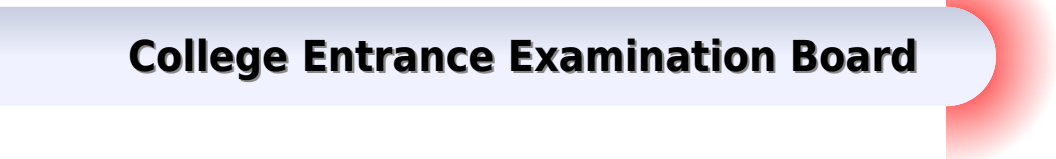
- 0.0010 M
 - 0.010 M
 - ☒ 0.10 M
 - 1.0 M

$$\frac{0.010 \text{ mol}}{0.1 \text{ L}} = 0.1 \text{ M}$$
- When Reaction 3 occurs, does the hybridization of the carbon atoms change?
 - ☒ Yes; it changes from sp to sp^2
 - Yes; it changes from sp to sp^3
 - Yes; it changes from sp^2 to sp^3
 - No; it does not change



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