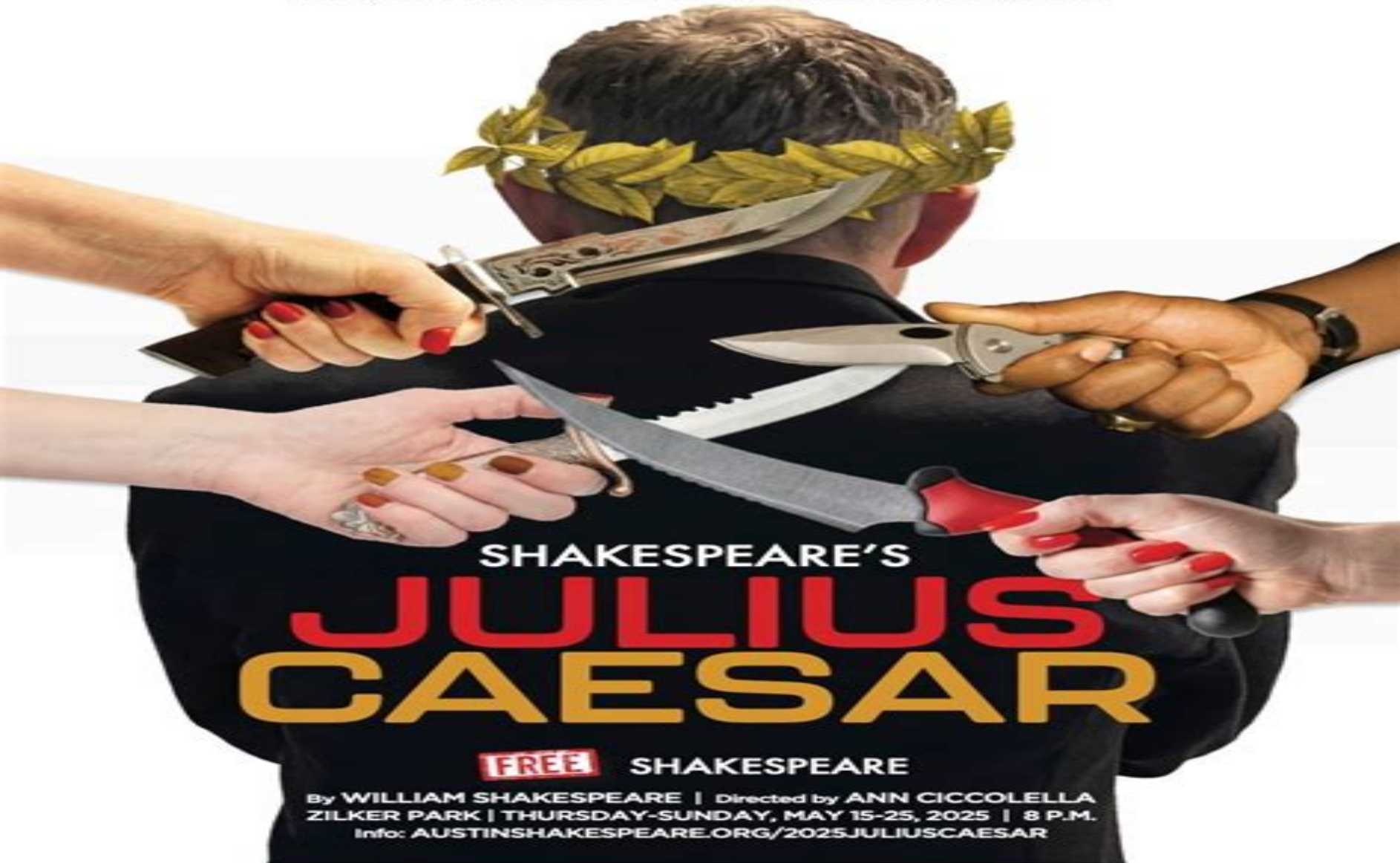


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xi PERSONAL NOTE FROM THE EDITOR
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geometries on paper with wedges and dashes to represent their 3d shape explain how π bonding affects the ability of atoms to rotate about a bond relate molecular polarity to physical properties

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web molecules composed of covalently bonded atoms may also be polar or nonpolar for the molecule to be polar it must of

course have polar bonds but the key factor for determining the polarity of a molecule is its shape if the polar bonds dipoles are symmetrical around the central atom they offset each other and the resulting molecule is

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web linear geometry all angles 180 degrees three regions of electron density around a central atom in a molecule form a

trigonal planar geometry all angles 120 degrees four regions of electron density around a central atom in a molecule form a

tetrahedral geometry all angles 109.5 degrees

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web in this investigation you will examine a theory that chemists use to explain different aspects of chemical bonding valence shell electron pair repulsion vsepr theory attention will be given to how molecules are arranged in different shapes and how chemists can predict the geometry of a given molecule