

# 2007 GMC Yukon Denali/Yukon XL Denali Owner Manual

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# 2007 Yukon Xl Denali Owners Manual

**S Ashworth**



## **2007 Yukon Xl Denali Owners Manual:**

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web we used mole fractions in gen chem 1 to describe gas mixtures where the partial pressure of a gas was its mole fraction

time the total pressure  $p_a$   $\times$   $p_t$  mole fraction  $a$  left  $\frac{p_a}{p_t}$  mole  $a$  moles total right  $\sum n$  1 all  $x_{n1}$  mole fraction of substance  $a$  is often denoted by  $x_a$  and

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web 1 in the given mixture the sum of all the mole fractions is equal to one  $x_a + x_b + \dots + x_n = 1$  2 the multiplication of the mole fraction by 100 gives the mole percentage 3 mole fraction is a unitless and dimensionless expression advantages of molar fraction mole fraction is not dependent on the temperature

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web nov 6 2023  $x_a$  mol  $a$  mol  $a$  mol  $b$  and  $x_b$  mol  $b$  mol  $a$  mol  $b$  note in the given mixture the sum of all the mole fractions present is equal to one  $x_a + x_b + \dots + x_n = 1$  when the mole fractions are multiplied by 100 they give the mole percentage mole fraction is a unitless and dimensionless expression image will be uploaded soon

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web you may work out the mole fractions on your own problem 5 what is the mass of a sample of a 0 449 molal  $KBr$  that

contains 2 92 kg of water 1 molality moles solute divided by kilograms solute 0 449 mol kg x 2 92 kg x 1 31108 mol of kbr 2 moles times molar mass equals grams 1 31108 mol 119 0023 g mol 156 g kbr

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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web since the mole fraction of the  $\text{O}_2$  is  $0.050$  it contributes  $0.050$  or if you will  $5.0\%$  of the total pressure thus total pressure is  $0.2105042 \text{ atm}$  problem 9 a sample of  $1.43 \text{ g}$  of helium and an unweighed quantity of  $\text{O}_2$  are mixed in a flask at room temperature