

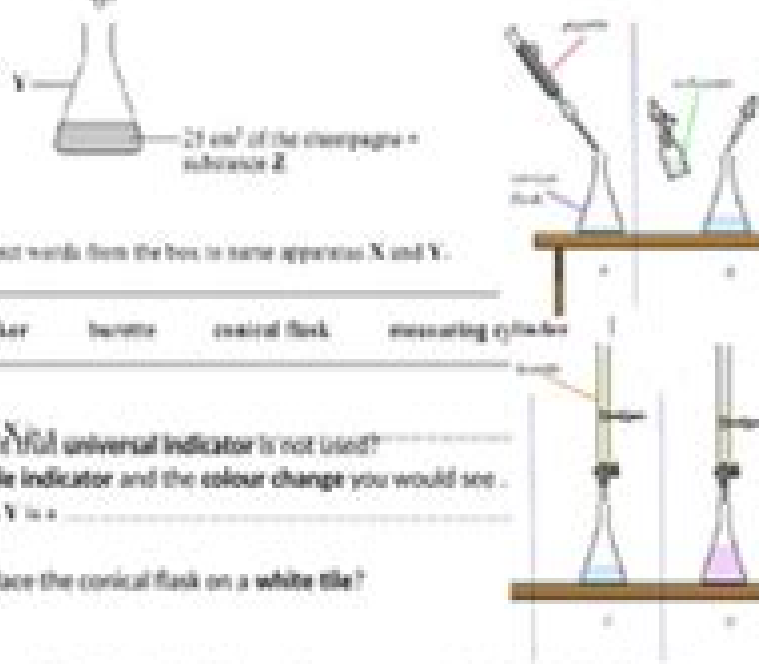
tration [Review]

Name: _____

Aim of practical: To determine the volume of sulfuric acid used to neutralise sodium hydroxide by titration

Key points: Writing the method and calculating concentration

a) Write out a **brief** method to determine the volume of sulfuric acid used to neutralise sodium hydroxide by titration and explain **why** you are doing each step. Use the diagrams to help you.



(a) Choose the correct words from the box to name apparatus X and Y.

beaker burette conical flask measuring cylinder

b) Why is a **universal indicator** not used?

Name a more suitable indicator and the colour change you would see.

(i) Apparatus Y is a _____

c) Why should you place the conical flask on a white tile?

d) Explain why a volumetric pipette should be used to measure the sodium hydroxide and a burette to measure the sulfuric acid.

Safety

a) For each of the following, give the hazard (the type of danger it poses to you)

- Dilute acid:

- Dilute alkali:

b) Give 2 safety precautions you should take when doing this practical:

Concordant titres

In titrations, only results 0.1 cm³ from each other are used to calculate the average titre. These are called **concordant** results.

Calculate the **average titre**

from the table results below:

	Titration 1	Titration 2	Titration 3	Titration 4	Titration 5
Volume of alkali used (cm ³)	13.50	12.10	11.10	12.10	12.10

Titration Calculations

- Use the volume and concentration of one reactant to calculate the moles.
- Use the chemical equation to find the moles of the other reactant.
- Calculate the volume or concentration as required of that reactant.

$$\text{concentration (mol/dm}^3\text{)} = \frac{\text{moles}}{\text{volume (dm}^3\text{)}}$$

To convert cm³ to dm³
divide by 1000.



To convert from mol/dm³ to g/dm³
times by the relative formula mass

a) 22.5cm³ of sodium hydroxide reacted with 25.0 cm³ of 0.500 mol/dm³ hydrochloric acid.



Calculate the concentration of the sodium hydroxide in mol/dm³.

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Chemistry: A Level Ted Lister, Janet Renshaw, 2016-05-05 Please note this title is suitable for any student studying Exam Board AQA Level A Level Subject Chemistry First teaching September 2015 First exams June 2017 Fully revised and updated for the new linear qualification written and checked by curriculum and specification experts this Student Book supports and extends students through the new course whilst delivering the maths practical and synoptic skills needed to succeed in the new A Levels and beyond The book uses clear straightforward explanations to develop real subject knowledge and allow students to link ideas together while developing essential exam skills **New 2015 A-Level Chemistry: AQA Year 1 & AS**

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