

1 Amended from AQA Chemistry Unit 2 Chemistry in Action CHEM2 January 2011
(Question 9)

- a** A student carried out a laboratory experiment to determine the enthalpy change when a sample of butan-1-ol was burnt.
The student found that the temperature of 175 g of water increased by 8.0 °C when 5.00×10^{-3} mol of pure butan-1-ol was burnt in air and the heat produced was used to warm the water.
Use the student's results to calculate a value, in kJ mol^{-1} , for the enthalpy change when one mole of butan-1-ol is burnt.
Give your answer to an appropriate number of significant figures.
(The specific heat capacity of water is $4.18 \text{ J K}^{-1} \text{ g}^{-1}$)

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(3 marks)

- b i** Give the meaning of the term *standard enthalpy of combustion*.

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(3 marks)

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D Kirk

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