

1 From AQA Chemistry Unit 1 Foundation Chemistry CHEM1 May 2014
(Question 1).

a Table 1.1 shows some data about fundamental particles in an atom.

Table 1.1 Mass of fundamental particles

Particle	proton	neutron	electron
Mass / g	1.6725×10^{-24}	1.6725×10^{-24}	0.0009×10^{-24}

- i An atom of hydrogen can be represented as ^1H . Use data from Table 1.1 to calculate the mass of this hydrogen atom.

$$1.6734 \times 10^{-27} \text{ kg}$$

(1 mark)

- ii Which one of the following is a fundamental particle that would not be deflected by an electric field?

- A electron
B neutron
C proton

Write the correct letter, A, B, or C, on the line below.

B

(1 mark)

- b A naturally occurring sample of the element boron has a relative atomic mass of 10.8. In this sample, boron exists as two isotopes, ^{10}B and ^{11}B .

- i Calculate the percentage abundance of ^{10}B in this naturally occurring sample of boron.

$$10.8 = (10)(x) + (11)(1-x)$$

$$10.8 = 10x + 11 - 11x$$

$$10.8 - 11 = 10x - 11x$$

$$-0.2 = -x$$

$$x = 0.2 \quad 0.2 \times 100 = 20\%$$

(2 marks)

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Guillaume Favre



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Table of Contents Aqa Science B1b 5 Exam Style Questions Answers

1. Understanding the eBook Aqa Science B1b 5 Exam Style Questions Answers
 - The Rise of Digital Reading Aqa Science B1b 5 Exam Style Questions Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Aqa Science B1b 5 Exam Style Questions Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Aqa Science B1b 5 Exam Style Questions Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Aqa Science B1b 5 Exam Style Questions Answers
 - Personalized Recommendations
 - Aqa Science B1b 5 Exam Style Questions Answers User Reviews and Ratings
 - Aqa Science B1b 5 Exam Style Questions Answers and Bestseller Lists

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 - Highlighting and Note-Taking Aqa Science B1b 5 Exam Style Questions Answers
 - Interactive Elements Aqa Science B1b 5 Exam Style Questions Answers
8. Staying Engaged with Aqa Science B1b 5 Exam Style Questions Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Aqa Science B1b 5 Exam Style Questions Answers
9. Balancing eBooks and Physical Books Aqa Science B1b 5 Exam Style Questions Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Aqa Science B1b 5 Exam Style Questions Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Aqa Science B1b 5 Exam Style Questions Answers
 - Setting Reading Goals Aqa Science B1b 5 Exam Style Questions Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Aqa Science B1b 5 Exam Style Questions Answers
 - Fact-Checking eBook Content of Aqa Science B1b 5 Exam Style Questions Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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