

Question number	Answer	Marks	Guidance
1 a	fizzes readily. X gets smaller and eventually dissolves completely X + sulfuric acid → X sulfate + hydrogen $X(s) + H_2SO_4(aq) \rightarrow XSO_4(aq) + H_2(g)$	1 1 1 2	
1 b	Magnesium reacts very slowly with water, calcium reacts vigorously with water. We do not know how much less reactive than calcium and more reactive than magnesium the element X is and so cannot predict whether the reaction will be slow like magnesium or more vigorous like calcium	1 1	
1 c i	the more reactive metal X displaces less reactive copper so solution loses its blue colour. X dissolves into solution and copper metal deposited as a brown solid $X(s) + Cu(NO_3)_2(aq) \rightarrow X(NO_3)_2(aq) + Cu(s)$	1 1 1	
1 c ii	$X(s) + Cu^{2+}(aq) \rightarrow X^{2+}(aq) + Cu(s)$	3	
1 c iii	$X(s) \rightarrow X^{2+}(aq) + 2e^-$ X atoms oxidised as lose 2 electrons $Cu^{2+}(aq) + 2e^- \rightarrow Cu(s)$ Cu^{2+} ions reduced as gain 2 electrons.	1 1 1 1	
1 d i	just above copper	1	
1 d ii	no change / no reaction as Y less reactive than magnesium so cannot displace it from solution of one of its compounds	1 1 1	
2 a i	react	1	
2 a ii	no reaction	1	

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Patrick Fullick



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