

Question number	Answer	Marks	Guidance
1 a i	Group 7	1	
1 a ii	Group 6	1	
1 a iii	Group 1	1	
GCSE Chemistry only			
1 a iv	between Groups 2 and 3	1	
1 b i	transition elements	1	
1 b ii	Group 1 / alkali metals	1	
1 b iii	Group 6 / noble gases	1	
1 b iv	Group 7 / halogens	1	
2 a	7 electrons as Group 7	1	
2 b	solid	1	
2 c i	ionic	1	
2 c ii	white	1	
2 c iii	NaAt	1	
2 c iv	sodium + astatine → sodium astatide $2\text{Na} + \text{At}_2 \rightarrow 2\text{NaAt}$	1	
2 c v	yes, chlorine more reactive than astatine so At displaced	1	

Question number	Answer	Marks	Guidance
2 a	F, Cl, Br, I, At The smaller the halogen atom, the more readily it accepts/ gains extra electron into its outermost shell, becomes halide ion with single negative charge, the stronger the electrostatic force of attraction between nucleus and extra electron entering smaller atom's outermost shell as electron closer to nuclear charge, and it becomes less shielded from nuclear charge than larger atoms with more inner shells of electrons. These two factors outweigh fact that nuclear charge is more positive going down Group 7.	1 1 1 1 1	
3 a	soft/good electrical conductivity low melting point	1 1 1	
3 b i	1+	1	
3 b ii	Rubidium compound rubidium iodide rubidium fluoride rubidium hydroxide	Chemical formula RbI RbF RbOH	3
3 c i	$2\text{Rb}(s) + 2\text{H}_2\text{O}(l) \rightarrow 2\text{RbOH}(aq) + \text{H}_2(g)$	3	
3 c ii	$2\text{Rb}(s) + \text{Cl}_2(g) \rightarrow 2\text{RbCl}(s)$	2	

Aqa Gcse Chemistry C2 Progress Check Answers

IM Harris



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Germany's answer to these British behemoths was the Albert Ballin designed trio of Imperator Vaterland and Bismarck. Through misfortune or war two of these vessels would sink but the others led useful lives with Aquitania surviving two world wars before being scrapped. Designed to be the absolute engineering achievements of their time, these nine vessels dominated the Atlantic. J Kent Layton tells the story of the Edwardian Superliners in this fabulously illustrated volume showcasing many images previously unpublished and never before seen. Rarely can one describe a book as definitive but this volume truly deserves the accolade. *AQA GCSE Chemistry*, 2011

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