

1 From AQA Chemistry Unit 1 Foundation Chemistry CHEM1 May 2014
(Question 1).

a Table 1.1 shows some data about fundamental particles in an atom.

Table 1.1 Mass of fundamental particles

Particle	proton	neutron	electron
Mass / g	1.6725×10^{-24}	1.6725×10^{-24}	0.0009×10^{-24}

- i** An atom of hydrogen can be represented as ^1H . Use data from Table 1.1 to calculate the mass of this hydrogen atom.

$$1.6734 \times 10^{-27} \text{ kg}$$

(1 mark)

- ii** Which one of the following is a fundamental particle that would not be deflected by an electric field?

- A** electron
B neutron
C proton

Write the correct letter, **A**, **B**, or **C**, on the line below.

B

(1 mark)

- b** A naturally occurring sample of the element boron has a relative atomic mass of 10.8. In this sample, boron exists as two isotopes, ^{10}B and ^{11}B .

- i** Calculate the percentage abundance of ^{10}B in this naturally occurring sample of boron.

$$10.8 = (10)(x) + (11)(1-x)$$

$$10.8 = 10x + 11 - 11x$$

$$10.8 - 11 = 10x - 11x$$

$$-0.2 = -x$$

$$x = 0.2 \quad 0.2 \times 100 = 20\%$$

(2 marks)

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Rob Ritchie



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