

1 From AQA Chemistry Unit 1 Foundation Chemistry CHEM1 May 2014
(Question 1).

a Table 1.1 shows some data about fundamental particles in an atom.

Table 1.1 Mass of fundamental particles

Particle	proton	neutron	electron
Mass / g	1.6725×10^{-24}	1.6725×10^{-24}	0.0009×10^{-24}

- i** An atom of hydrogen can be represented as ${}^1\text{H}$. Use data from Table 1.1 to calculate the mass of this hydrogen atom.

$$1.6734 \times 10^{-27} \text{ kg}$$

(1 mark)

- ii** Which one of the following is a fundamental particle that would not be deflected by an electric field?

- A** electron
B neutron
C proton

Write the correct letter, **A**, **B**, or **C**, on the line below.

B

(1 mark)

- b** A naturally occurring sample of the element boron has a relative atomic mass of 10.8. In this sample, boron exists as two isotopes, ${}^{10}\text{B}$ and ${}^{11}\text{B}$.

- i** Calculate the percentage abundance of ${}^{10}\text{B}$ in this naturally occurring sample of boron.

$$10.8 = (10)(x) + (11)(1-x)$$

$$10.8 = 10x + 11 - 11x$$

$$10.8 - 11 = 10x - 11x$$

$$-0.2 = -x$$

$$x = 0.2 \quad 0.2 \times 100 = 20\%$$

(2 marks)

Aqa Chemistry A Level Exam Style Questions Answers

Chapter 11

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decomposition of group II carbonates thermal decomposition of group II nitrates uses of group II elements uses of group II metals uses of halogens and their compounds The Halogenoalkanes Quiz Questions PDF e Book Chapter 14 interview questions and answers on Halogenoalkanes uses of halogenoalkanes elimination reactions nucleophilic substitution in halogenoalkanes and nucleophilic substitution reactions The Hydrocarbons Quiz Questions PDF e Book Chapter 15 interview questions and answers on Introduction to alkanes sources of alkanes addition reactions of alkenes alkane reaction alkenes and formulas The Introduction to Organic Chemistry Quiz Questions PDF e Book Chapter 16 interview questions and answers on Organic chemistry functional groups organic reactions naming organic compounds stereoisomerism structural isomerism and types of organic reactions The Ionic Equilibria Quiz Questions PDF e Book Chapter 17 interview questions and answers on Introduction to ionic equilibria buffer solutions equilibrium and solubility indicators and acid base titrations pH calculations and weak acids The Lattice Energy Quiz Questions PDF e Book Chapter 18 interview questions and answers on Introduction to lattice energy ion polarization lattice energy value atomization and electron affinity Born Haber cycle and enthalpy changes in solution The Moles and Equations Quiz Questions PDF e Book Chapter 19 interview questions and answers on Amount of substance atoms molecules mass chemical formula and equations gas volumes mole calculations relative atomic mass solutions and concentrations The Nitrogen and Sulfur Quiz Questions PDF e Book Chapter 20 interview questions and answers on Nitrogen gas nitrogen and its compounds nitrogen and gas properties ammonia ammonium compounds environmental problems caused by nitrogen compounds and nitrate fertilizers sulfur and oxides sulfuric acid and properties and uses of sulfuric acid The Organic and Nitrogen Compounds Quiz Questions PDF e Book Chapter 21 interview questions and answers on Amides in chemistry amines amino acids peptides and proteins The Periodicity Quiz Questions PDF e Book Chapter 22 interview questions and answers on Acidic oxides basic oxides aluminum oxide balancing equation period 3 chlorides balancing equations reactions with chlorine balancing equations reactions with oxygen bonding nature of period 3 oxides chemical properties of chlorine chemical properties of oxygen chemical properties periodicity chemistry periodic table chemistry oxides chlorides of period 3 elements electrical conductivity in period 3 oxides electronegativity of period 3 oxides ionic bonds molecular structures of period 3 oxides oxidation number of oxides oxidation numbers oxides and hydroxides of period 3 elements oxides of period 3 elements period III chlorides periodic table electronegativity physical properties periodicity reaction of sodium and magnesium with water and relative melting point of period 3 oxides The Polymerization Quiz Questions PDF e Book Chapter 23 interview questions and answers on Types of polymerization polyamides polyesters and polymer deductions The Rates of Reaction Quiz Questions PDF e Book Chapter 24 interview questions and answers on Catalysis collision theory effect of concentration reaction kinetics and temperature effect on reaction rate The Reaction Kinetics Quiz Questions PDF e Book Chapter 25 interview questions and answers on Reaction kinetics catalysts kinetics and reaction mechanism order of reaction rate constant k and rate of reaction The Redox Reactions

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elements physical properties of group II elements physical properties of group VII elements reaction of group II elements with oxygen reactions of group II elements reactions of group VII elements thermal decomposition of carbonates and nitrates thermal decomposition of group II carbonates thermal decomposition of group II nitrates uses of group II elements uses of group II metals uses of halogens and their compounds The Halogenoalkanes MCQ PDF e Book Chapter 14 practice test to solve MCQ questions on Halogenoalkanes uses of halogenoalkanes elimination reactions nucleophilic substitution in halogenoalkanes and nucleophilic substitution reactions The Hydrocarbons MCQ PDF e Book Chapter 15 practice test to solve MCQ questions on Introduction to alkanes sources of alkanes addition reactions of alkenes alkane reaction alkenes and formulas The Introduction to Organic Chemistry MCQ PDF e Book Chapter 16 practice test to solve MCQ questions on Organic chemistry functional groups organic reactions naming organic compounds stereoisomerism structural isomerism and types of organic reactions The Ionic Equilibria MCQ PDF e Book Chapter 17 practice test to solve MCQ questions on Introduction to ionic equilibria buffer solutions equilibrium and solubility indicators and acid base titrations pH calculations and weak acids The Lattice Energy MCQ PDF e Book Chapter 18 practice test to solve MCQ questions on Introduction to lattice energy ion polarization lattice energy value atomization and electron affinity Born Haber cycle and enthalpy changes in solution The Moles and Equations MCQ PDF e Book Chapter 19 practice test to solve MCQ questions on Amount of substance atoms molecules mass chemical formula and equations gas volumes mole calculations relative atomic mass solutions and concentrations The Nitrogen and Sulfur MCQ PDF e Book Chapter 20 practice test to solve MCQ questions on Nitrogen gas nitrogen and its compounds nitrogen and gas properties ammonia ammonium compounds environmental problems caused by nitrogen compounds and nitrate fertilizers sulfur and oxides sulfuric acid and properties and uses of sulfuric acid The Organic and Nitrogen Compounds MCQ PDF e Book Chapter 21 practice test to solve MCQ questions on Amides in chemistry amines amino acids peptides and proteins The Periodicity MCQ PDF e Book Chapter 22 practice test to solve MCQ questions on Acidic oxides basic oxides aluminum oxide balancing equation period 3 chlorides balancing equations reactions with chlorine balancing equations reactions with oxygen bonding nature of period 3 oxides chemical properties of chlorine chemical properties of oxygen chemical properties periodicity chemistry periodic table chemistry oxides chlorides of period 3 elements electrical conductivity in period 3 oxides electronegativity of period 3 oxides ionic bonds molecular structures of period 3 oxides oxidation number of oxides oxidation numbers oxides and hydroxides of period 3 elements oxides of period 3 elements period III chlorides periodic table electronegativity physical properties periodicity reaction of sodium and magnesium with water and relative melting point of period 3 oxides The Polymerization MCQ PDF e Book Chapter 23 practice test to solve MCQ questions on Types of polymerization polyamides polyesters and polymer deductions The Rates of Reaction MCQ PDF e Book Chapter 24 practice test to solve MCQ questions on Catalysis collision theory effect of concentration reaction kinetics and temperature effect on reaction rate The Reaction Kinetics MCQ PDF e

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Table of Contents Aqa Chemistry A Level Exam Style Questions Answers Chapter 11

1. Understanding the eBook Aqa Chemistry A Level Exam Style Questions Answers Chapter 11
 - The Rise of Digital Reading Aqa Chemistry A Level Exam Style Questions Answers Chapter 11
 - Advantages of eBooks Over Traditional Books
2. Identifying Aqa Chemistry A Level Exam Style Questions Answers Chapter 11
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Aqa Chemistry A Level Exam Style Questions Answers Chapter 11
 - User-Friendly Interface
4. Exploring eBook Recommendations from Aqa Chemistry A Level Exam Style Questions Answers Chapter 11
 - Personalized Recommendations

- Aqa Chemistry A Level Exam Style Questions Answers Chapter 11 User Reviews and Ratings
- Aqa Chemistry A Level Exam Style Questions Answers Chapter 11 and Bestseller Lists
- 5. Accessing Aqa Chemistry A Level Exam Style Questions Answers Chapter 11 Free and Paid eBooks
 - Aqa Chemistry A Level Exam Style Questions Answers Chapter 11 Public Domain eBooks
 - Aqa Chemistry A Level Exam Style Questions Answers Chapter 11 eBook Subscription Services
 - Aqa Chemistry A Level Exam Style Questions Answers Chapter 11 Budget-Friendly Options
- 6. Navigating Aqa Chemistry A Level Exam Style Questions Answers Chapter 11 eBook Formats
 - ePub, PDF, MOBI, and More
 - Aqa Chemistry A Level Exam Style Questions Answers Chapter 11 Compatibility with Devices
 - Aqa Chemistry A Level Exam Style Questions Answers Chapter 11 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Aqa Chemistry A Level Exam Style Questions Answers Chapter 11
 - Highlighting and Note-Taking Aqa Chemistry A Level Exam Style Questions Answers Chapter 11
 - Interactive Elements Aqa Chemistry A Level Exam Style Questions Answers Chapter 11
- 8. Staying Engaged with Aqa Chemistry A Level Exam Style Questions Answers Chapter 11
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Aqa Chemistry A Level Exam Style Questions Answers Chapter 11
- 9. Balancing eBooks and Physical Books Aqa Chemistry A Level Exam Style Questions Answers Chapter 11
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Aqa Chemistry A Level Exam Style Questions Answers Chapter 11
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Aqa Chemistry A Level Exam Style Questions Answers Chapter 11
 - Setting Reading Goals Aqa Chemistry A Level Exam Style Questions Answers Chapter 11
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Aqa Chemistry A Level Exam Style Questions Answers Chapter 11
 - Fact-Checking eBook Content of Aqa Chemistry A Level Exam Style Questions Answers Chapter 11

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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