- Amended from AQA Chemistry Unit 2 Chemistry in Action CHEM2 January 2011 (Question 9)
 - a A student carried out a laboratory experiment to determine the enthalpy change when a sample of butan-1-ol was burnt.

The student found that the temperature of 175 g of water increased by 8.0 °C when 5.00×10^{-3} mol of pure butan-1-ol was burnt in air and the heat produced was used to warm the water.

Use the student's results to calculate a value, in kJ mol⁻¹, for the enthalpy change when one mole of butan-1-ol is burnt.

Give your answer to an appropriate number of significant figures.

(The specific heat capacity of water is 4.18 J K⁻¹ g⁻¹)

	Give the	meaning:	of the term	standard.	enthaloy of	combustion

(3 marks)

(3 marks)

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ML Yell

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