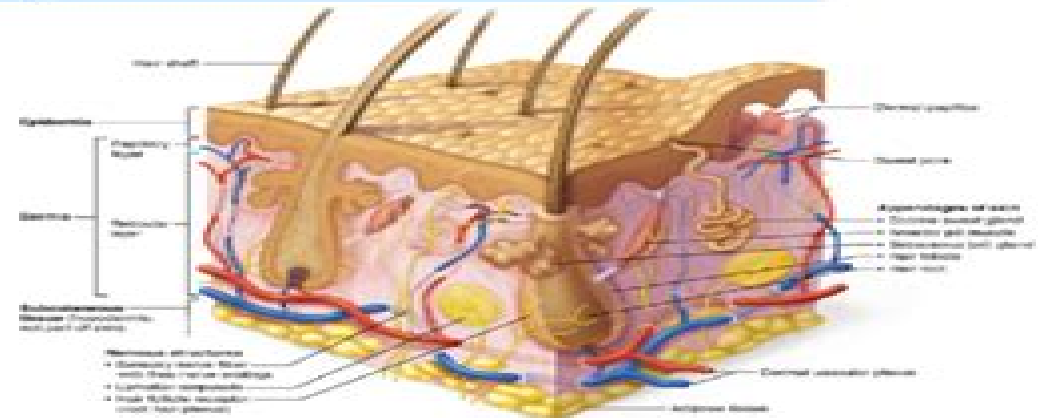


CHAPTER 5

The Integumentary System

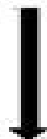
5.1 – THE SKIN CONSISTS OF 2 LAYERS: THE EPIDERMIS AND DERMIS

- 2 layers:
 1. **Epidermis:** Composed of epithelial cells; outermost protective shield of the body
 - Epi = Upon
 2. **Dermis:** Underlying bulk of skin; tough, leathery layer composed mostly of dense connective tissue
 - Highly vascularized
- Nutrients reach epidermis by diffusing through tissue fluid via BV in the dermis
- **Hypodermis (Subcutaneous) Tissue:** Subcutaneous tissue deep to skin
 - Not part of skin; also known as superficial fascia
 - Stores adipose tissue; shock absorber and insulator (↓ heat loss)
 - Anchors the skin to the underlying structures
- Sliding skin protects by ensuring many blows bounce off the skin



5.2 – THE EPIDERMIS IS A KERATINIZED STRATIFIED SQUAMOUS EPITHELIUM

SUPERFICIAL



DEEP

- There are 4 or 5 layers of the epidermis:
 - (I) Stratum Corneum
 - (II) Stratum Lucideum
 - (III) Stratum Granulosum
 - (IV) Stratum Spinosum
 - (V) Stratum Basale
- 4 distinct cells:
 - (I) Keratinocytes
 - (II) Melanocytes
 - (III) Dendritic (Langerhans) Cells
 - (IV) Tactile (Merkel) Cells

Anatomy Packet Answers For Integumentary System

Brigitte Niedzwiecki, Julie Pepper



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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction $2\text{Ag}^+ \text{aq}$ $\text{Cr}_2\text{O}_7^{2-} \text{aq}$ H_2O $2\text{Cr}^{3+} \text{aq}$ $7\text{H}^+ \text{aq}$

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web chemical equation in which all dissolved ionic reactants and products including spectator ions are explicitly represented by formulas for their dissociated ions molecular equation chemical equation in which all reactants and products are

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web jan 29 2020 the ionic form of this reaction is $2\text{Na}^+\text{aq} + 2\text{Cl}^-\text{aq} + \text{Cu}^{2+}\text{aq} + \text{SO}_4^{2-}\text{aq} \rightarrow 2\text{Na}^+\text{aq} + \text{SO}_4^{2-}\text{aq} + \text{CuCl}_2\text{s}$ the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web mar 11 2014 solution ionic equation $\text{H}^+\text{aq} + \text{Cl}^-\text{aq} + \text{Na}^+\text{aq} + \text{HCO}_3^-\text{aq} \rightarrow \text{Na}^+\text{aq} + \text{Cl}^-\text{aq} + \text{H}_2\text{O}\text{l} + \text{CO}_2\text{g}$ the spectator ions are Cl^- and Na^+
net ionic equation $\text{H}^+\text{aq} + \text{HCO}_3^-\text{aq} \rightarrow \text{H}_2\text{O}\text{l} + \text{CO}_2\text{g}$ here is a video on spectator ions
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web the six Na^+aq ions and the six Cl^-aq ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3\text{Ba}^{2+}\text{aq} + 2\text{PO}_4^{3-}\text{aq} \rightarrow \text{Ba}_3(\text{PO}_4)_2\text{s}$

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web these spectator ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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 $\text{AgNO}_3\text{aq} + \text{KCl}\text{aq} \rightarrow \text{AgCl}\text{s} + \text{KNO}_3\text{aq}$ total ionic $\text{Ag}^+\text{aq} + \text{NO}_3^-\text{aq} + \text{K}^+\text{aq} + \text{Cl}^-\text{aq} \rightarrow \text{AgCl}\text{s} + \text{K}^+\text{aq} + \text{NO}_3^-\text{aq}$ net ionic $\text{Ag}^+\text{aq} + \text{Cl}^-\text{aq} \rightarrow \text{AgCl}\text{s}$ 2

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