

Experiment 10: Electrodes with a Standard Solution

From: *Basic Laboratory Experiments*, Experiment 10, 45-50

The purpose of this experiment is to determine the standard potential for several cells consisting of $\text{Zn} / \text{Zn}^{2+}$ and $\text{Cu} / \text{Cu}^{2+}$ electrodes. Electrode potentials are a tool to support calculations such as those involving Gibbs free energy, Gibbs free energy of activation, and other thermodynamic quantities. The standard potential for a half-cell is the potential of the half-cell when the concentration of the ions is 1.0 M. The standard potential for a half-cell is the potential of the half-cell when the concentration of the ions is 1.0 M. The standard potential for a half-cell is the potential of the half-cell when the concentration of the ions is 1.0 M. The standard potential for a half-cell is the potential of the half-cell when the concentration of the ions is 1.0 M.

In this experiment, you will determine the standard potential for several half-cells and calculate the standard potential for the overall cell reaction.

Objectives:

- to determine the standard potential for several half-cells and calculate the standard potential for the overall cell reaction
- to calculate the standard potential for the overall cell reaction
- to determine the standard potential for the overall cell reaction
- to determine the standard potential for the overall cell reaction
- to calculate the standard potential for the overall cell reaction

Materials:

- 1.0 M Zn^{2+} solution
- 1.0 M Cu^{2+} solution
- 1.0 M Zn^{2+} solution
- 1.0 M Cu^{2+} solution
- 1.0 M Zn^{2+} solution
- 1.0 M Cu^{2+} solution
- 1.0 M Zn^{2+} solution
- 1.0 M Cu^{2+} solution
- 1.0 M Zn^{2+} solution
- 1.0 M Cu^{2+} solution

Procedure:

- 1.0 M Zn^{2+} solution
- 1.0 M Cu^{2+} solution
- 1.0 M Zn^{2+} solution
- 1.0 M Cu^{2+} solution
- 1.0 M Zn^{2+} solution
- 1.0 M Cu^{2+} solution
- 1.0 M Zn^{2+} solution
- 1.0 M Cu^{2+} solution
- 1.0 M Zn^{2+} solution
- 1.0 M Cu^{2+} solution

Results:

- 1.0 M Zn^{2+} solution
- 1.0 M Cu^{2+} solution
- 1.0 M Zn^{2+} solution
- 1.0 M Cu^{2+} solution
- 1.0 M Zn^{2+} solution
- 1.0 M Cu^{2+} solution
- 1.0 M Zn^{2+} solution
- 1.0 M Cu^{2+} solution
- 1.0 M Zn^{2+} solution
- 1.0 M Cu^{2+} solution

Discussion:

- 1.0 M Zn^{2+} solution
- 1.0 M Cu^{2+} solution
- 1.0 M Zn^{2+} solution
- 1.0 M Cu^{2+} solution
- 1.0 M Zn^{2+} solution
- 1.0 M Cu^{2+} solution
- 1.0 M Zn^{2+} solution
- 1.0 M Cu^{2+} solution
- 1.0 M Zn^{2+} solution
- 1.0 M Cu^{2+} solution

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Lingsheng Yao



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