

End of chapter questions

Getting started

- 1 Identify the type of bonding in potassium bromide, KBr.
a metallic b ionic c covalent d giant 1 Mark
- 2 Which substance is made of small molecules?
a carbon dioxide b silicon dioxide c magnesium oxide d sodium sulfide 1 Mark
- 3 Use two 'particle' diagrams to show the differences between a liquid and a gas. 2 Marks
- 4 Substance R conducts electricity when solid and is malleable. What is the structure of R?
a simple molecular b giant ionic c giant covalent d giant metallic 1 Mark
- 5 Substance G has a melting point of $-33\text{ }^{\circ}\text{C}$ and a boiling point of $52\text{ }^{\circ}\text{C}$. Explain which state it is in at $20\text{ }^{\circ}\text{C}$. 1 Mark
- 6 Lithium chloride, LiCl, is made of ions. Explain if it will conduct electricity. 2 Marks
- 7 Match the symbol to the type of particle. 2 Marks

K^+	small molecule
NO_2	polymer
$-(\text{XZ})_n$	ion

Going further

- 8 Describe the structure of diamond. 1 Mark
- 9 Suggest one structural feature that graphene and fullerenes have in common. 1 Mark
- 10 Cupronickel is an alloy made from copper and nickel. Describe two ways in which the properties of the alloy differ from the individual metals. 2 Marks
- 11 Determine the structure of A and B. Justify your answer. 4 Marks

Substance	A	B
Mp and Bp / $^{\circ}\text{C}$	-205 and -192	-420 and 907
Conducts electricity?	No	Yes when solid

- 12 Metal J has a melting point of $232\text{ }^{\circ}\text{C}$ and metal L has a melting point of $328\text{ }^{\circ}\text{C}$. When they are mixed together the mixture has a lower melting point than either of these two temperatures. Explain what is being made and why its properties make it useful. 2 Marks

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Lyn Nicholls, Ken Gadd



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