

- 5 Limestone contains calcium carbonate (CaCO_3).
- 5 (a) Name the type of reaction that takes place when calcium carbonate is heated strongly.
Name the products formed.
- [3 marks]

- 5 (b) A student investigated what happens when limestone is heated strongly.

This is the method the student used:

- measure the mass of limestone before heating
- measure the mass of solid product after heating
- repeat the experiment three more times.

The student's results are shown in Table 3.

Table 3

	Experiment 1	Experiment 2	Experiment 3	Experiment 4
Mass of limestone before heating in g	5.0	5.0	5.0	5.0
Mass of solid product after heating in g	3.2	4.0	3.1	3.3
Mass lost in g	1.8	1.0	1.9	1.7

- 5 (b) (i) Calculate the mean mass lost, taking account of any anomalies.
- [2 marks]

- 5 (b) (ii) The student used the chemical equation to calculate the maximum mass lost by 5.0 g of calcium carbonate when heated.

The maximum mass lost is 2.2 g.

Suggest and explain **two** reasons why the mean mass lost in the experiments to heat limestone is less than 2.2 g.

Do **not** include any reference to weighing errors in your answer.

[4 marks]

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**Nick Dixon, Nick England, Richard
Grime, Nora Henry, Ali Hodgson, Steve
Witney**

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