

CHAPTER 10: Reactions

Reaction Energy**SECTION 1**

Write a balanced chemical equation for the following reactions in the space provided.

- For elements in their standard state, the value of ΔH_f° is _____.
- The formation and decomposition of a substance is represented by the following thermodynamic equations:



- Is the energy being added to or is it being subtracted from the decomposition?
- What is the correct sign for the enthalpy change in this decomposition reaction?

Write the correct answer in the space to the left. Show all your work in the space provided.

- _____ $\Delta H_{\text{f}}^\circ$ for a reaction is 2076 kJ whereas a 4 mol/L of energy is released from the decomposition. Is it? _____
- _____ A compound has a specific heat of $0.001 \text{ J/g}^\circ\text{C}$. If a 10 g sample undergoes a 1°C increase in temperature, how much energy is released? _____
- The products is a molecular form of energy of 100 g/mol , and the reactant has an energy of 100 g/mol . _____
- _____ What is the value of ΔH for this reaction? _____

Answers For Reaction Energy Section 1

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Answers For Reaction Energy Section 1:

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