# Stoichiometry: Mass - Mole Worksheet

Based on the following chemical equation, please answer the following questions!

Fe (SO<sub>4</sub>), + 6 NaCl --> 2 FeCl, + 3 Na<sub>2</sub>SO<sub>4</sub>

Calculate the amount of moles of NaCl needed to produce 75 g of FeCl,.

9 FeCl,	1 mole FeCl,	mole NaCl	
	9 FeCl,	mole FeCl	

Calculate the amount of Fe, (SO,), in grams produce 215 g Na, SO,.

9 Na <sub>3</sub> SO <sub>4</sub>	1 mole Na SO	mole Fe <sub>2</sub> (SO <sub>4</sub> ),	9 Fe <sub>2</sub> (SO <sub>4</sub> ),
	9 Na <sub>2</sub> SO <sub>4</sub>	mole Na,SO,	1 mole Fe <sub>2</sub> (SO <sub>4</sub> ),

Calculate the amount in grams of FeCl, produced with 95g of Fe (SO<sub>4</sub>),

9 Fe <sub>2</sub> (SO <sub>4</sub> ),	mole Fe2(SO4)	mole FeCl,	9 FeCl,
	9 Fe,(SO,),	mole Fe(SO <sub>4</sub> ),	1 mole FeCl,

Calculate the amount of moles of Fe<sub>2</sub>(SO<sub>4</sub>) needed to react with 123.5 g of NaCl

9 NaCl	1 mole NaCl	mole Fe(SO <sub>e</sub> ),	
	9 NaCl	mole NaCl	

Based on the following chemical equation, please answer the following questions!

9 Ca(HCO <sub>y</sub> ) <sub>2</sub>	1 mole Ca(HCO <sub>5</sub> ),	mole CaCO,
	9 Ca(HCO,),	mole Ca(HCO,),

# **Answers For Study Guide Mass Moles**

**Susan McMurry** 

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