

Kinetics Worksheet 1

3. For the reaction: $A + 2B \rightarrow 3C + D$
the following data were obtained,

run	[A]	[B]	rate $M \cdot s^{-1}$
1	0.16	0.36	2.67
2	0.16	0.09	0.67
3	0.32	0.24	7.11
4	0.08	0.09	0.168
5	0.24	0.12	2.0
6	0.02	1.35	?

$$R = k[A]^y[B]^x$$

- (A) Write the rate law expression for this reaction.

(B) $\left(\frac{0.36}{0.09}\right)^x = \frac{2.67 \text{ m s}^{-1}}{0.67 \text{ m s}^{-1}}$ (A) $\left(\frac{0.16}{0.08}\right)^y = \frac{0.67 \text{ m s}^{-1}}{0.168 \text{ m s}^{-1}}$
 $(4)^x = 4$ $(2)^y = 4$
 $x \ln 4 = \ln 4$ $y \ln 2 = \ln 4$
 $x = 1$ $y = 2$

$$R = k[A]^2[B]$$

- (B) Calculate the value of k, the rate constant (remember units).

$$k = \frac{R}{[A]^2[B]} = \frac{2.67 \text{ m s}^{-1}}{(0.16)^2 [0.36]} \quad k = \text{m}^{-2} \text{s}^{-1}$$

$$k = 290 \text{ m}^{-2} \text{s}^{-1}$$

- (C) Calculate the rate of the reaction when $[A] = 0.02 \text{ M}$ and $[B] = 1.35 \text{ M}$.

$$R = 290 \text{ m}^{-2} \text{s}^{-1} [0.02]^2 [1.35] = 0.16 \frac{\text{m}}{\text{s}}$$

- (D) Assume that the reaction goes to completion. Under the conditions specified in the second run, what would the final [C] be?



$$\text{limit} \Rightarrow 0.09 B \times \frac{1A}{2B} = 0.045 A$$

$$0.045 A \times \frac{3C}{1A} = 0.135 \text{ m C}$$

$$\begin{aligned} .16 A \times \frac{3C}{1A} &= .48 \text{ m} \\ .09 B \times \frac{3C}{2B} &= .135 \text{ m} \end{aligned}$$

or R

Answers Kinetics Worksheet 1

Sir Tarun Rupani



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