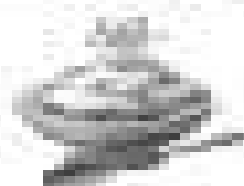


Atomic Probing Structure of the Atom

Name _____
Date _____



Purpose:

To determine the structure of the atom by observing the deflection of alpha particles.

Part I: Understanding the Structure of the Atom

The Rutherford experiment was a landmark in the history of physics. It showed that the atom is not a uniform sphere of positive charge, as was previously believed, but that it has a small, dense, positively charged nucleus. The results of the experiment are summarized in the following table.

Observation: Most alpha particles pass straight through the foil without deflection.

Conclusion: Most of the atom is empty space. The positive charge is concentrated in a small nucleus.

Observation: Some alpha particles are deflected at large angles.

Conclusion: The nucleus is small, dense, and positively charged. It contains most of the mass of the atom.

Observation: A few alpha particles are deflected at very large angles.

Conclusion: The nucleus is very small and contains most of the mass of the atom.

Answers Lesson 11 Atomic Pudding

RS Peters



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Decoding **Answers Lesson 11 Atomic Pudding**: Revealing the Captivating Potential of Verbal Expression

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