

Writing Linear Equations

Name _____

Class Period _____

Write an equation of the line that contains the given point and has the given slope.

1. $(-7, 5)$, slope = -2

$$\begin{aligned} y &= mx + b \\ 5 &= -2(-7) + b \\ 5 &= +14 + b \\ -14 & \\ -9 &= b \end{aligned}$$

$$y = -2x - 9$$

2. $(-4, -4)$, slope = -1

$$\begin{aligned} -4 &= -1(-4) + b \\ -4 &= 4 + b \\ -4 & -4 \\ -8 &= b \end{aligned}$$

$$y = -x - 8$$

3. $(-4, -7)$, slope = -2

$$\begin{aligned} -7 &= -2(-4) + b \\ -7 &= 8 + b \\ b &= -15 \end{aligned}$$

$$y = -2x - 15$$

4. $(-1, -1)$, slope = 4

$$\begin{aligned} -1 &= 4(-1) + b \\ -1 &= -4 + b \\ b &= 3 \end{aligned}$$

$$y = 4x + 3$$

5. $(-4, -15)$, slope = -5

$$\begin{aligned} -15 &= -5(-4) + b \\ -15 &= 20 + b \\ -20 & -20 \\ -35 &= b \end{aligned}$$

$$y = -5x - 35$$

6. $(-8, 12)$, slope = $1/2$

$$\begin{aligned} 12 &= 1/2(-8) + b \\ 12 &= -4 + b \\ b &= 16 \end{aligned}$$

$$y = 1/2x + 16$$

7. $(-7, 7)$, slope = -5

$$\begin{aligned} 7 &= -5(-7) + b \\ 7 &= 35 + b \\ b &= -28 \end{aligned}$$

$$y = -5x - 28$$

8. $(8, 4)$, slope = 0

$$y = 4$$

9. $(6, -2)$, slope = undefined

$$x = 6$$

10. $(8, -4)$, slope = $-1/4$

$$\begin{aligned} -4 &= -1/4(8) + b \\ -4 &= -2 + b \\ b &= -2 \end{aligned}$$

$$y = -1/4x - 2$$

11. $(-3, -6)$, slope = $1/3$

$$\begin{aligned} -6 &= 1/3(-3) + b \\ -6 &= -1 + b \\ b &= -5 \end{aligned}$$

$$y = 1/3x - 5$$

12. $(-4, 2)$, slope = -2

$$\begin{aligned} 2 &= -2(-4) + b \\ 2 &= 8 + b \\ b &= -6 \end{aligned}$$

$$y = -2x - 6$$

Algebra January 2014 Answer Key

Mark Steyvers

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web in today's experiment you will determine the solubility product ksp of calcium hydroxide Ca(OH)_2 by measuring the concentration of Ca(OH)_2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of Ca(OH)_2 is represented in equation form as shown below $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ the solubility product expression describes in mathematical terms the equilibrium

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web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ $5 \times 10^{-2} \text{ M}$ calcium ions $5 \times 10^{-2} \text{ M}$ $2 \times 10^{-2} \text{ M}$ hydroxide ions calculate calcium hydroxide ksp using the concentrations of calcium ions and hydroxide ions

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web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the ksp of a slightly soluble basic salt and how solubility changes with temperature ksp is another form of equilibrium expression that deals with solubility the reaction that will take place is $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$

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molarity 0.050 M volume of filtered Ca(OH)_2 solution mL 5.00 5.00 3.71 13.75 equivalence point volume mL OH⁻ in saturated Ca(OH)_2 M Ca^{2+} in saturated Ca(OH)_2 M K_{sp} of Ca(OH)_2 average K_{sp} of Ca(OH)_2

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web the chemical equation used was $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ the solubility product was determined to be K_{sp} Ca^{2+} OH⁻ from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

18.1 solubility product constant K_{sp} chemistry libretexts - Feb 01 2023

web Jul 12 2023 because the concentration of a pure solid such as $\text{Ca}_3(\text{PO}_4)_2$ is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $K_{\text{Ca}_3(\text{PO}_4)_2} = [\text{Ca}^{2+}]^3 [\text{PO}_4^{3-}]^2$

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web lab report 8 determining of calcium hydroxide the purpose of this was to titrate a saturated Ca(OH)_2 solution with a standard HCl solution and find the K_{sp} or the OH⁻ concentration was found and then the concentration was found and used for this experiment a titration of calcium hydroxide with hydrochloric acid

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