



General Chemistry II Equation Sheet

Think about how to set up the problem first, then apply the needed principles and formulas.

Phase Changes

$$\Delta T_b = iK_b m$$

$$\Delta T_f = iK_f m$$

$$\pi = iMRT$$

$$c = kP$$

$$\frac{P_1}{P_2} = \frac{Z_1}{Z_2}$$

$$P_{\text{solution}} = \chi_{\text{solvent}} P_{\text{solvent}}^{\circ}$$

$$\Delta P = P_{\text{solvent}}^{\circ} - P_{\text{solution}} = \chi_{\text{solute}} P_{\text{solvent}}^{\circ}$$

Cells

$$E = E^{\circ} - \frac{RT}{nF} \ln(Q)$$

$$E = E^{\circ} - \frac{0.0592}{n} \log(Q) \quad (\text{at } 25^{\circ}\text{C})$$

$$\Delta G = -nFE_{\text{cell}}$$

$$\Delta G^{\circ} = -nFE_{\text{cell}}^{\circ}$$

$$E_{\text{cell}}^{\circ} = \frac{RT}{nF} \ln(K)$$

$$E_{\text{cell}}^{\circ} = \frac{0.0592\text{V}}{n} \log(K) \quad (\text{at } 25^{\circ}\text{C})$$

USE ONE:

$$E_{\text{cell}}^{\circ} = E_{\text{cathode(re.d.)}}^{\circ} - E_{\text{anode(re.d.)}}^{\circ}$$

$$E_{\text{cell}}^{\circ} = E_{\text{cathode(re.d.)}}^{\circ} + E_{\text{anode(ox.)}}^{\circ}$$

Thermodynamics

$$\Delta G = \Delta H - T\Delta S$$

$$\Delta S_{\text{surroundings}} = \frac{-\Delta H_{\text{sys}}}{T}$$

$$\Delta S_{\text{universe}} = \Delta S_{\text{sys}} + \Delta S_{\text{surroundings}}$$

$$\Delta S_{\text{rxn}}^{\circ} = \sum S_{\text{products}}^{\circ} - \sum S_{\text{reactants}}^{\circ}$$

$$\Delta G_{\text{rxn}}^{\circ} = \sum \Delta G_{\text{products}}^{\circ} - \sum \Delta G_{\text{reactants}}^{\circ}$$

Concentrations

$$\text{molality} = \frac{\text{moles solute}}{\text{mass solvent (kg)}}$$

$$\% \text{ by mass} = \frac{\text{mass solute}}{\text{mass solute} + \text{mass solvent}} \times 100\%$$

$$\text{parts per million} = \frac{\text{mass solute}}{\text{mass solute} + \text{mass solvent}} \times 10^6$$

$$\text{parts per billion} = \frac{\text{mass solute}}{\text{mass solute} + \text{mass solvent}} \times 10^9$$

Acids and Bases

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = K_a \times K_b$$

$$\text{pOH} = -\log[\text{OH}^-]$$

$$[\text{OH}^-] = 10^{-\text{pOH}}$$

$$\text{pH} + \text{pOH} = \text{p}K_a + \text{p}K_b = 14.00 \quad (\text{at } 25^{\circ}\text{C})$$

$$\text{pH} = \text{p}K_a + \log \frac{[\text{A}^-]}{[\text{HA}]}$$

$$\text{pOH} = \text{p}K_b + \log \frac{[\text{BH}^+]}{[\text{B}]}$$

$$\% \text{ ionization} = \frac{[\text{H}_3\text{O}^+]_{\text{eq}}}{[\text{HA}]_{\text{eq}}} \times 100\%$$

$$\text{p}K_a = -\log(K_a)$$

$$\text{p}K_b = -\log(K_b)$$

$$K_a = 10^{-\text{p}K_a}$$

$$K_b = 10^{-\text{p}K_b}$$

Equilibria

$$Q_c = \frac{[\text{C}]^c[\text{D}]^d}{[\text{A}]^a[\text{B}]^b}$$

$$K_c = \frac{[\text{C}]^c[\text{D}]^d}{[\text{A}]^a[\text{B}]^b}$$

$$Q_p = \frac{P_{\text{C}}^c P_{\text{D}}^d}{P_{\text{A}}^a P_{\text{B}}^b}$$

$$K_p = \frac{P_{\text{C}}^c P_{\text{D}}^d}{P_{\text{A}}^a P_{\text{B}}^b}$$

$$K_p = K_c [R \times T]^{\Delta n}$$

$$\Delta G = \Delta G^{\circ} + RT \ln(Q)$$

$$\Delta G^{\circ} = -RT \ln(K)$$

$$\text{rate} = k[\text{A}]^x[\text{B}]^y$$

$$k = Ae^{-\frac{E_a}{RT}}$$

$$\ln\left(\frac{k_2}{k_1}\right) = \frac{E_a}{R} \left(\frac{1}{T_1} - \frac{1}{T_2}\right)$$

$$\text{half life} = \frac{\ln(2)}{k} \quad (\text{1st order reactions only})$$

$$[\text{A}]_t = -kt + [\text{A}]_0$$

$$\ln([\text{A}]_t) = -kt + \ln([\text{A}]_0)$$

$$\frac{1}{[\text{A}]_t} = kt + \frac{1}{[\text{A}]_0}$$

Answer Sheet General Chemistry Equations

Andrei Hent, John Andraos



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General Chemistry Donald A. McQuarrie, Stanley Gill, 2011-06-15 This Fourth Edition of McQuarrie's classic text offers a thorough revision and a quantum leap forward from the previous edition Taking an atoms first approach it promises to be another ground breaking text in the tradition of McQuarrie's many previous works This outstanding new text available in a soft cover edition offers professors a fresh choice and outstanding value

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Table of Contents Answer Sheet General Chemistry Equations

1. Understanding the eBook Answer Sheet General Chemistry Equations
 - The Rise of Digital Reading Answer Sheet General Chemistry Equations
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Sheet General Chemistry Equations
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Sheet General Chemistry Equations
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Sheet General Chemistry Equations
 - Personalized Recommendations
 - Answer Sheet General Chemistry Equations User Reviews and Ratings
 - Answer Sheet General Chemistry Equations and Bestseller Lists

5. Accessing Answer Sheet General Chemistry Equations Free and Paid eBooks
 - Answer Sheet General Chemistry Equations Public Domain eBooks
 - Answer Sheet General Chemistry Equations eBook Subscription Services
 - Answer Sheet General Chemistry Equations Budget-Friendly Options
6. Navigating Answer Sheet General Chemistry Equations eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Sheet General Chemistry Equations Compatibility with Devices
 - Answer Sheet General Chemistry Equations Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Sheet General Chemistry Equations
 - Highlighting and Note-Taking Answer Sheet General Chemistry Equations
 - Interactive Elements Answer Sheet General Chemistry Equations
8. Staying Engaged with Answer Sheet General Chemistry Equations
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Sheet General Chemistry Equations
9. Balancing eBooks and Physical Books Answer Sheet General Chemistry Equations
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Sheet General Chemistry Equations
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer Sheet General Chemistry Equations
 - Setting Reading Goals Answer Sheet General Chemistry Equations
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Sheet General Chemistry Equations
 - Fact-Checking eBook Content of Answer Sheet General Chemistry Equations
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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