

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive the empirical formula of a pure substance and various unknowns forming multicomponent systems.
2. To calculate the number of moles of each of the multicomponent system in the system.
3. To determine the molar fraction of each.
4. To determine the theoretical amount of various substances that would be produced.
5. To compare the theoretical amount to the actual amount of substances formed and calculate the percent yield.

Basic calculations:

Weight of Sodium Chloride solution: 25.0 ml

Weight of Calcium Chloride solution: 25.0 ml

Weight of Dry Filter Paper: 0.00 g

Weight of Dry Filter Paper + Dry Solid: 0.00 g

Describe what happens when mixing the two solutions:

Initially the solution turned white and cloudy. After stirring with a bar magnet, the solution became white and the bottom portion turned black.

the white precipitate indicates the reaction product that formed in the

Results:

Quantities and Calculations:

(a) Calculations:

$$\text{moles of sodium chloride} = 0.025 \text{ L} \times 0.1 \text{ M} = \frac{0.0025 \text{ mol}}{1000 \text{ mL}} \times 58.44 \text{ g/mol} = 0.000146 \text{ mol}$$

$$\text{moles of calcium chloride} = 0.025 \text{ L} \times 0.1 \text{ M} = \frac{0.0025 \text{ mol}}{1000 \text{ mL}} \times 146.01 \text{ g/mol} = 0.000365 \text{ mol}$$

$$\text{mass of sodium chloride produced} = 0.0025 \text{ L} \times 0.1 \text{ M} \times 58.44 \text{ g/mol} = 0.0146 \text{ g}$$

$$\text{mass of calcium chloride produced} = 0.025 \text{ L} \times 0.1 \text{ M} \times 146.01 \text{ g/mol} = 0.365 \text{ g}$$

(b) Chemical equation: $\text{Na}_2\text{CO}_3 + \text{CaCl}_2 \rightarrow \text{CaCO}_3 + 2\text{NaCl}$

(c) Reaction product:

Calcium carbonate

Since Na_2CO_3 is 0.0025 mol and CaCl_2 is 0.0025 mol, the limiting reagent is

Na_2CO_3 and the limiting reagent is 0.0025 mol

$$\text{mass of } \text{CaCO}_3 \text{ formed using } \text{Na}_2\text{CO}_3 = 0.0025 \text{ mol} \times \frac{100.09 \text{ g/mol}}{1000 \text{ mL}} = 0.250 \text{ g}$$

Answers For Lab 6b Heath Chemistry

Sebastian Brünink



Answers For Lab 6b Heath Chemistry:

Phlebotomy Essentials Ruth E. McCall, 2023-08-11 Phlebotomy Essentials Eighth Edition provides accurate up to date and practical information and instruction in phlebotomy procedures and techniques along with a comprehensive background in phlebotomy theory and principles It is appropriate for use as an instructional text or as a reference for those who wish to update skills or study for national certification This updated edition details how today s phlebotomists work in an approach optimized for how today s students learn Combined with an optional Workbook Exam Review book and updated digital courseware the latest edition of Phlebotomy Essentials represents a cornerstone of preparation for a successful career in phlebotomy Complete program integration of textbook workbook and exam review products working together as a complete learning system for optimal student outcomes Phlebotomy Essentials Eighth Edition with Navigate Premier Access Student Workbook for Phlebotomy Essentials Eighth Edition Phlebotomy Exam Review Eighth Edition with Navigate Premier Access Phlebotomy Essentials Updated to current CLSI guidelines TestPrep alerts throughout the textbook share current statistics on which questions students typically get wrong and why Map of NAACLS Standards where each competency is covered Certification Test Plan Alignment Map including American Medical Technologists AMT Registered Phlebotomy Technician American Society for Clinical Pathology ASCP Phlebotomy Technician National Healthcareer Association NHA Certified Phlebotomy Technician National Center for Competency Testing NCCT National Certified Phlebotomy Technician 2024 516 pages **Whitaker's Cumulative Book List** ,1972 **El-Hi Textbooks in Print** ,1984 **Comprehensive Organic Chemistry Experiments for the Laboratory Classroom** Carlos A. M. Afonso, Nuno R. Candeias, Dulce Pereira Simão, Alexandre F. Trindade, Jaime A. S. Coelho, Bin Tan, Robert Franzén, 2016-12-16 This expansive and practical textbook contains organic chemistry experiments for teaching in the laboratory at the undergraduate level covering a range of functional group transformations and key organic reactions The editorial team have collected contributions from around the world and standardized them for publication Each experiment will explore a modern chemistry scenario such as sustainable chemistry application in the pharmaceutical industry catalysis and material sciences to name a few All the experiments will be complemented with a set of questions to challenge the students and a section for the instructors concerning the results obtained and advice on getting the best outcome from the experiment A section covering practical aspects with tips and advice for the instructors together with the results obtained in the laboratory by students has been compiled for each experiment Targeted at professors and lecturers in chemistry this useful text will provide up to date experiments putting the science into context for the students *Chemist and Druggist* ,1869 *Books in Print* ,1994 *Catalog of Copyright Entries. Third Series* Library of Congress. Copyright Office, 1961 Includes Part 1 Number 1 2 Books and Pamphlets Including Serials and Contributions to Periodicals January December *The British National Bibliography* Arthur James Wells, 1965
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