

Experiment 2 - Gravimetric Analysis of a Chloride Salt (G)

Purpose:

Determining quantitative value of chloride in an unknown soluble salt by showing and explaining common methods in gravimetric analysis.

Theory:

- Reaction between silver ion and chloride ion to form silver chloride as an equation:



- AgCl solubility in water and the K_{sp} :

Silver chloride is very insoluble, so we can assume that the above reaction goes to completion.

$$K_{sp} = 1.6 \times 10^{-10}$$

- AgCl is precipitated with an acidic solution to reduce likelihood of acid ions interfering with results. These ions will form co-precipitates with silver in the acid, due to these co-precipitates, the results would be higher.
- To prevent the precipitate from going through the filtered paper, it is heated to coagulate it into a different texture that's easier to work with. Heating the precipitate makes it coagulate, therefore, it is harder for the precipitate to go through the filtered paper.
- Dry silver chloride decomposes into silver and chlorine in the presence of light:



- If photodecomposition takes place in air, results will be lower, since chlorine will be released in the form of gas. If photochemical decomposition occurs in the presence of excess silver ions in aqueous solution, it will result in higher results.

Analysis Of An Unknown Salt Lab Report

RJ Shavelson



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Laboratory Manual for Principles of General Chemistry Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures Laboratory Experiments for General Chemistry Harold R. Hunt, Toby F. Block, 1994 **English in**

Analytical Chemistry. Communicating about Methods & Techniques. Книга для студента Надежда

Зорина, Александра Соболева, 2022-10-19

Scientific and Technical Aerospace Reports, 1995-05

General

Chemistry Hunt, 1976-06-07

Laboratory Manual for Principles of General Chemistry J. A. Beran, Mark

Lassiter, 2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section **General, Organic, and Biological Chemistry Study Guide and Selected Solutions** Karen C. Timberlake, 2001-11 Keyed to the learning goals in the text this guide is designed to promote active learning through a variety of exercises with answers and mastery exams The guide also contains complete solutions to odd numbered problems

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Fundamentals of Inorganic Chemistry Bruce E. Wilcox,1997-07 Chemistry Karen Timberlake,1999 Suitable for one or two term lab courses covering general organic and biological chemistry this new edition written by Karen Timberlake features many improvements to the insightful experiments that have made it the leading lab manual Each experiment encourages critical thinking with laboratory goals discussion of related concepts clear instructions new pre lab questions and comprehensive report pages Forty one experiments illustrate the basic principles of chemistry

Report summaries
United States. Environmental Protection Agency,1983 Science Educator's Guide to Laboratory Assessment Rodney L. Doran,2002 The book opens with an up to date discussion of assessment theory research and uses Then comes a wealth of sample assessment activities in biology chemistry physics and Earth science Keyed to the National Science Education

Standards the activities include reproducible task sheets and scoring rubrics All are ideal for helping students reflect on their own learning during science lab Nuclear Science Abstracts ,1964 *U.S. Government Research & Development Reports* ,1968 Government Reports Announcements & Index ,1996 **Bibliography of Scientific and Industrial Reports** ,1948

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