

6. Figure 11-4 is a diagram of the frontal section of the heart. Follow the instructions below to complete this exercise.

First, draw arrows to indicate the direction of blood flow through the heart. Draw the pathway of the oxygen-rich blood with red arrows, and trace the pathway of oxygen-poor blood with blue arrows.

Second, identify each of the elements of the intrinsic conduction system (numbers 1-5 on the figure) by inserting the appropriate terms in the blanks left of the figure. Then, indicate with green arrows the pathway that impulses take through this system.

Third, correctly identify each of the heart valves (numbers 6-9 on the figure) by inserting the appropriate terms in the blanks left of the figure, and draw in and identify by name the cordlike structures that anchor the flaps of the atrioventricular (AV) valves.

Fourth, use the numbers from the figure to identify the structures described below. Place the numbers in the lettered answer blanks.

- | | | | | |
|----------|----|----------|----|---|
| <u>6</u> | A. | <u>7</u> | B. | Prevent backflow into the ventricles when the heart is relaxed |
| <u>8</u> | C. | <u>9</u> | D. | Prevent backflow into the atria when the ventricles are contracting |
| <u>9</u> | E. | | | AV valve with three flaps |
| <u>8</u> | F. | | | AV valve with two flaps |
| <u>1</u> | G. | | | The pacemaker of the intrinsic conduction system |
| <u>2</u> | H. | | | The point in the intrinsic conduction system where the impulse is temporarily delayed |

Figure 11-4



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Robyn Beirman



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web mar 6 2017 1 the chemical reaction $2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$ is a synthesis reaction decomposition reaction single displacement reaction double displacement reaction combustion reaction you can recognize a decomposition or analysis reaction by noticing a compound or molecule breaks into smaller chemical species 2 the chemical reaction

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web answer chemical kinetics is the branch of physical chemistry that is concerned with understanding the rates of chemical reactions it is to be contrasted with thermodynamics which deals with the direction in which a process occurs but in

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web 1 the chemical equations in model 1 contain the phase notations s l g and aq write each symbol by each meaning dissolved in water liquid solid gas aq l s g 2 based on the

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combustion reaction predict the products of simple reactions the chemical reactions we have described are only a tiny sampling of the infinite number of chemical reactions possible

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web an answer key document has been included and is available for download some additional information to be aware of for teachers demonstrations 1 and 5 will not produce a chemical reaction five of the reactions will produce indicators of a chemical reaction for students to observe

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web identifying chemical reaction types key part 1 identify the type of reaction 1 $\text{Na}_3\text{PO}_4 + 3\text{KOH} \rightarrow 3\text{NaOH} + \text{K}_3\text{PO}_4$ double displacement 2 $\text{MgCl}_2 + \text{Li}_2\text{CO}_3 \rightarrow \text{MgCO}_3 + 2\text{LiCl}$ double displacement 3 $\text{C}_6\text{H}_{12} + 9\text{O}_2 \rightarrow 6\text{CO}_2 + 6\text{H}_2\text{O}$ combustion 4 $\text{Pb} + \text{FeSO}_4 \rightarrow \text{PbSO}_4 + \text{Fe}$ single displacement 5 $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$ decomposition 6 $\text{P}_4 + 3\text{O}_2 \rightarrow 2\text{P}_2\text{O}_3$

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web 1 3nabr 1h3po4 1na3po4 3hbr type of reaction double displacement 2 3ca oh 2 1al2 so4 33caso4 2al oh 3 type of reaction
double displacement 3 3mg 1fe2o3 2fe 3mgo type of reaction single displacement 4 1c2h4 3o2 2co2 2h2o type of reaction
combustion 5 2pbso4 2pbso3 1o2 type of reaction

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starting substances are called the reactants and the new substances that form are called the products

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differentiate between a physical and chemical change physical does not change the identity of the matter can be changed
back easily chemical new substance is created with new physical and chemical