

SCH4U Unit 4 - Quiz – Equilibrium

Player: _____

Period: _____

Q1: (10)

Q2: (10)

Q3: (10)

Q4: (10)

Q5: (10)

Write the formulas for:

$$K_{eq} = \frac{[products]}{[reactants]} \quad ; \quad K_{eq} = \frac{P_{products}}{P_{reactants}}$$

$$K_{eq} = \frac{P_{CO_2}^2 P_{H_2O}}{P_{CH_4}^2 P_{O_2}^2} \quad ; \quad K_{eq} = \frac{P_{CO_2}^2 P_{H_2O}}{(P_{CH_4})^2 (P_{O_2})^2}$$

gpt (aq) + h₂o (l) ⇌ gpt(s) + h₂g (g) ; gpt(s) ⇌ h₂g (g)

K_{eq} = K_{sp}/K_a

Question 1:

1. Hydrogen sulfide, H₂S, is a slightly poisonous gas. At 25°C, an equilibrium mixture contains 0.025 moles of hydrogen, H₂, 0.025 moles of hydrogen sulfide, with an equilibrium pressure of 0.400 atm. The reaction is:



If the volume of an 8.00 L container at 25°C, where a concentration of H₂ is present at equilibrium:

- What is the concentration of H₂ at equilibrium?
- Calculate the value of K_{eq} at equilibrium.
- Calculate the value of K_{eq} at equilibrium.
- What other equilibria have chemical formula involving H₂ and hydrogen sulfide?
- Calculate:

	2H ₂ (g)	H ₂ (g)	H _{2S} (g)
Initial (moles) (moles)			
Change (moles) (moles)			
Final (moles) (moles)	0.025	0.025	0

For the following equilibrium:

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BrightThinker



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