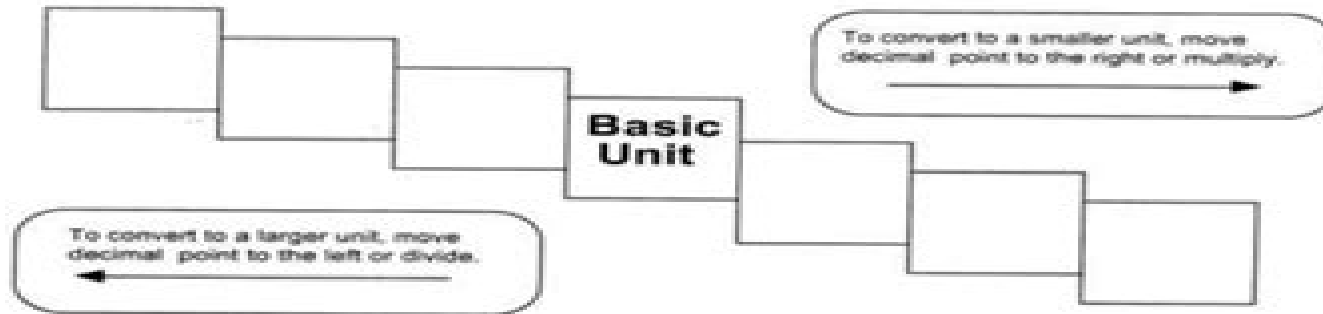


Metric Mania Conversion Practice

Name _____



Try these conversions, using the ladder method.

$$1000 \text{ mg} = \mathbf{1.000} \text{ g}$$

$$1 \text{ L} = \mathbf{1000} \text{ mL}$$

$$160 \text{ cm} = \mathbf{1600} \text{ mm}$$

$$14 \text{ km} = \mathbf{14000} \text{ m}$$

$$109 \text{ g} = \mathbf{0.109} \text{ kg}$$

$$250 \text{ m} = \mathbf{0.250} \text{ km}$$

Compare using <, >, or =.

$$56 \text{ cm} \mathbf{<} 6 \text{ m}$$

$$7 \text{ g} \mathbf{>} 698 \text{ mg}$$

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Puja Mehta



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Close Reading with Paired Texts Level 5: Engaging Lessons to Improve Comprehension Oczkus, Lori, 2017-03-01 Prepare fifth grade students for college and career readiness with this content packed resource Authored by Lori Oczkus and Timothy Rasinski this resource includes 12 units across the four content areas of language arts science social studies and mathematics Each unit incorporates close reading paired fiction and nonfiction text passages text dependent questions comparing and contrasting text and hands on activities to unify each week s worth of lessons Differentiation and reciprocal teaching strategies and assessment options are also included within each unit to tailor to multiple intelligences and monitor students progress

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$$\begin{aligned} & \text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3 \\ & 2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2 \\ & 2\text{NaF} + \text{Cl}_2 \rightarrow 2\text{H}_2 + \text{O}_2 \\ & 2\text{H}_2\text{O} + \text{Pb} + \text{OH}^- \rightarrow 2\text{HCl} + 2\text{H}_2\text{O} + 1\text{PbCl}_2 \\ & \text{AlBr}_3 + 3\text{K}_2\text{SO}_4 \rightarrow 6\text{KBr} + \text{Al}_2\text{SO}_4 + 3\text{CH}_4 + 2\text{O}_2 \\ & \text{CO}_2 + 2\text{H}_2\text{O} + \text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O} + \text{C}_8\text{H}_{18} + 25\text{O}_2 \\ & \text{CO}_2 + 16\text{CO}_2 + 18\text{H}_2\text{O} \end{aligned}$$

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 hen balancing

$$\begin{aligned} & 1\text{Na} + 3\text{PO}_4 + 3\text{KOH} + 3\text{NaOH} + 1\text{K} + 3\text{PO}_4 + 1\text{MgF}_2 + 1\text{Li} + 2\text{CO}_3 + 1\text{MgCO}_3 + 2\text{LiF} + 1\text{P}_4 + 3\text{O}_2 + 2\text{P}_2 + \text{O}_3 + 2\text{RbNO}_3 + 1\text{BeF}_2 + 1\text{BeNO}_3 + 2\text{RbF} \end{aligned}$$

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 $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$
 $2\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF} + \text{Cl}_2$
 $2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$
 $\text{Pb} + \text{HCl} \rightarrow \text{PbCl}_2 + \text{H}_2$
 $2\text{Al} + 3\text{K}_2\text{SO}_4 \rightarrow 6\text{KBr} + \text{Al}_2\text{SO}_4 + 3\text{H}_2$
 $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
 $8\text{H}_2\text{O} + \text{C}_3\text{H}_8 \rightarrow 5\text{O}_2 + 3\text{CO}_2 + 4\text{H}_2$

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 $1\text{ N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$
 $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$
 $2\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF} + \text{Cl}_2$
 $2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$
 $\text{Pb} + \text{HCl} \rightarrow \text{PbCl}_2 + \text{H}_2$
 $2\text{Al} + 3\text{K}_2\text{SO}_4 \rightarrow 6\text{KBr} + \text{Al}_2\text{SO}_4 + 3\text{H}_2$
 $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$

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 $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$
 $2\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF} + \text{Cl}_2$
 $2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$
 $\text{Pb} + \text{HCl} \rightarrow \text{PbCl}_2 + \text{H}_2$
 $2\text{Al} + 3\text{K}_2\text{SO}_4 \rightarrow 6\text{KBr} + \text{Al}_2\text{SO}_4 + 3\text{H}_2$
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 $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
 $2\text{Fe} + 3\text{H}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2$
 $\text{C}_4\text{H}_{10} + 6\text{O}_2 \rightarrow 4\text{CO}_2 + 5\text{H}_2\text{O}$

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 $3\text{Fe} + 2\text{O}_3 \rightarrow 3\text{FeO}$
 $2\text{Fe} + 3\text{CO} \rightarrow 2\text{Fe}_3\text{O}_4 + 3\text{H}_2$
 $3\text{BaO} + 3\text{H}_2\text{O} \rightarrow 4\text{Mg} + 4\text{NH}_3$
 $4\text{Mg} + 2\text{P} \rightarrow 7\text{N}_2 + 3\text{H}_2 + \text{O}_2$
 $4\text{Al} + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Al}_2\text{O}_3 + 3\text{H}_2$
 $4\text{Al} + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Al}_2\text{O}_3 + 3\text{H}_2$
 $4\text{Al} + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Al}_2\text{O}_3 + 3\text{H}_2$
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