

Types of Chemical Reaction Worksheet

Balance the reactions **1 to 6** and indicate which type of chemical reaction (synthesis, decomposition, single-displacement, double-displacement or combustion) is being represented:

- | | |
|---|-------------------------------------|
| 1. $2 \text{NaBr} + \text{Ca(OH)}_2 \rightarrow \text{CaBr}_2 + 2 \text{NaOH}$ | Reaction Type : single displacement |
| 2. $2 \text{NH}_3 + \text{H}_2\text{SO}_4 \rightarrow 3 (\text{NH}_4)_2\text{SO}_4$ | Reaction Type : synthesis |
| 3. $4 \text{C}_3\text{H}_8\text{O} + 27 \text{O}_2 \rightarrow 20 \text{CO}_2 + 18 \text{H}_2\text{O}$ | Reaction Type : combustion |
| 4. $3 \text{Pb} + 2 \text{H}_3\text{PO}_4 \rightarrow 3 \text{H}_2 + \text{Pb}_3(\text{PO}_4)_2$ | Reaction Type : single displacement |
| 5. $\text{Li}_2\text{N} + 3 \text{NH}_4\text{NO}_3 \rightarrow 3 \text{LiNO}_3 + (\text{NH}_4)_3\text{N}$ | Reaction Type : double displacement |
| 6. $\text{HBr} + 3 \text{Al(OH)}_3 \rightarrow \text{H}_2\text{O} + 3 \text{AlBr}_3$ | Reaction Type : double displacement |

Indicate which type of chemical reaction (synthesis, decomposition, single-displacement, double-displacement or combustion) is being represented in 7 to 20.

- | | |
|--|--|
| 7. $\text{Na}_3\text{PO}_4 + 3 \text{KOH} \rightarrow 3 \text{NaOH} + \text{K}_3\text{PO}_4$ | Reaction Type double displacement |
| 8. $\text{MgCl}_2 + \text{Li}_2\text{CO}_3 \rightarrow \text{MgCO}_3 + 2 \text{LiCl}$ | Reaction Type double displacement |
| 9. $\text{C}_8\text{H}_{12} + 9 \text{O}_2 \rightarrow 6 \text{CO}_2 + 6 \text{H}_2\text{O}$ | Reaction Type combustion |
| 10. $\text{Pb} + \text{FeSO}_4 \rightarrow \text{PbSO}_4 + \text{Fe}$ | Reaction Type single displacement |
| 11. $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$ | Reaction Type decomposition |
| 12. $\text{P}_4 + 3 \text{O}_2 \rightarrow 2 \text{P}_2\text{O}_3$ | Reaction Type synthesis |
| 13. $2 \text{RbNO}_3 + \text{BeF}_2 \rightarrow \text{Be(NO}_3)_2 + 2 \text{RbF}$ | Reaction Type double displacement |
| 14. $2 \text{AgNO}_3 + \text{Cu} \rightarrow \text{Cu(NO}_3)_2 + 2 \text{Ag}$ | Reaction Type single displacement |
| 15. $\text{C}_3\text{H}_8\text{O} + 4 \text{O}_2 \rightarrow 3 \text{CO}_2 + 3 \text{H}_2\text{O}$ | Reaction Type combustion |
| 16. $2 \text{C}_2\text{H}_5 + \text{Fe} \rightarrow \text{Fe(C}_2\text{H}_5)_2$ | Reaction Type synthesis |
| 17. $\text{SeCl}_6 + \text{O}_2 \rightarrow \text{SeO}_2 + 3 \text{Cl}_2$ | Reaction Type single displacement |
| 18. $2 \text{MgI}_2 + \text{Mn(SO}_4)_2 \rightarrow 2 \text{MgSO}_4 + \text{MnI}_4$ | Reaction Type double displacement |
| 19. $\text{O}_3 \rightarrow \text{O} + \text{O}_2$ | Reaction Type decomposition |
| 20. $2 \text{NO}_2 \rightarrow 2 \text{O}_2 + \text{N}_2$ | Reaction Type decomposition |

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