

Stoichiometry Worksheet #2 (mole-mass, mass-mole problems)



- a. If 15.0g of N_2O_4 was produced, how many moles of O_2 were required?

$$\frac{15.0\text{g N}_2\text{O}_4}{92.0\text{g N}_2\text{O}_4} \times \frac{1\text{ mol N}_2\text{O}_4}{1\text{ mol N}_2\text{O}_4} \times \frac{2\text{ mol O}_2}{1\text{ mol N}_2\text{O}_4} = 0.326\text{ mol O}_2$$

- b. If 4.0×10^{-3} moles of oxygen reacted, how many grams of N_2 were needed?

$$\frac{4.0 \times 10^{-3}\text{ mol O}_2}{2\text{ mol O}_2} \times \frac{1\text{ mol N}_2}{1\text{ mol O}_2} \times \frac{28.0\text{g N}_2}{1\text{ mol N}_2} = 5.6 \times 10^{-2}\text{ g N}_2$$



- a. What is the mass of potassium nitrate that is produced when 2.04 moles of potassium phosphate react?

$$\frac{2.04\text{ mol K}_3\text{PO}_4}{1\text{ mol K}_3\text{PO}_4} \times \frac{3\text{ mol KNO}_3}{1\text{ mol K}_3\text{PO}_4} \times \frac{101.1\text{ g KNO}_3}{1\text{ mol KNO}_3} = 619\text{g KNO}_3$$

- b. If 5.80g of aluminum phosphate are formed, how many moles of aluminum nitrate reacted?

$$\frac{5.80\text{g AlPO}_4}{122.0\text{g AlPO}_4} \times \frac{1\text{ mol AlPO}_4}{1\text{ mol AlPO}_4} \times \frac{1\text{ mol Al}(\text{NO}_3)_3}{1\text{ mol AlPO}_4} = 4.75 \times 10^{-2}\text{ mol Al}(\text{NO}_3)_3$$



- a. If you have 5.50 mol of CaC_2 , how much C_2H_2 do you get?

$$\frac{5.50\text{ mol CaC}_2}{1\text{ mol CaC}_2} \times \frac{1\text{ mol C}_2\text{H}_2}{1\text{ mol CaC}_2} \times \frac{26.0\text{g C}_2\text{H}_2}{1\text{ mol C}_2\text{H}_2} = 143\text{g C}_2\text{H}_2$$

- b. How many moles of water are needed when 65.0g of CaC_2 have reacted?

$$\frac{65.0\text{g CaC}_2}{64.1\text{g CaC}_2} \times \frac{1\text{ mol CaC}_2}{1\text{ mol CaC}_2} \times \frac{2\text{ mol H}_2\text{O}}{1\text{ mol CaC}_2} = 2.03\text{ mol H}_2\text{O}$$

4. In photosynthesis, water reacts with carbon dioxide to give oxygen and glucose ($\text{C}_6\text{H}_{12}\text{O}_6$). Write and balance the chemical equation. How many moles of CO_2 are required to make 120.0g of glucose?



$$\frac{120.0\text{g C}_6\text{H}_{12}\text{O}_6}{180.0\text{g C}_6\text{H}_{12}\text{O}_6} \times \frac{1\text{ mol C}_6\text{H}_{12}\text{O}_6}{1\text{ mol C}_6\text{H}_{12}\text{O}_6} \times \frac{6\text{ mol CO}_2}{1\text{ mol C}_6\text{H}_{12}\text{O}_6} = 4.000\text{ mol CO}_2$$

Answers For The Chemistry Lab Of Stoichiometry

Joseph A. Mascetta, Mark Kernion



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Educational Series, Mark Kernion, Joseph A. Mascetta, 2021-09-07 provides a complete guide to the fundamentals of chemistry Page 4 of cover

Multiple Solution Methods for Teaching Science in the Classroom Stephen DeMeo, 2008 For the first time in science education the subject of multiple solution methods is explored in book form While a multiple method teaching approach is utilized extensively in math education there are very few journal articles and no texts written on this topic in science Teaching multiple methods to science students in order to solve quantitative word problems is important for two reasons First it challenges the practice by teachers that one specific method should be used when solving problems Secondly it calls into question the belief that multiple methods would confuse students and retard their learning Using a case study approach and informed by research conducted by the author this book claims that providing students with a choice of methods as well as requiring additional methods as a way to validate results can be beneficial to student learning A close reading of the literature reveals that time spent on elucidating concepts rather than on algorithmic methodologies is a critical issue when trying to have students solve problems with understanding It is argued that conceptual understanding can be enhanced through the use of multiple methods in an environment where students can compare evaluate and verbally discuss competing methodologies through the facilitation of the instructor This book focuses on two very useful methods proportional reasoning PR and dimensional analysis DA These two methods are important because they can be used to solve a large number of problems in all of the four academic sciences biology chemistry physics and earth science This book concludes with a plan to integrate DA and PR into the academic science curriculum starting in late elementary school through to the introductory college level A challenge is presented to teachers as well as to textbook writers who rely on the single method paradigm to consider an alternative way to teach scientific problem solving

Cooperative Chemistry Lab Manual Cooper, 2005-02 The laboratory course described in the lab manual emphasizes experimental design data analysis and problem solving Inherent in the design is the emphasis on communication skills both written and oral Students work in groups on open ended projects in which they are given an initial scenario and then asked to investigate a problem There are no formalized instructions and students must plan and carry out their own investigations

Chemistry: The Easy Way Joseph A. Mascetta, Mark Kernion, 2019-08-06 A self teaching guide for students Chemistry The Easy Way provides easy to follow lessons with comprehensive review and practice This edition features a brand new design and new content structure with illustrations and practice questions An essential resource for High school and college courses Virtual learning Learning pods Homeschooling Chemistry The Easy Way covers Atomic Structure Chemical Formulas Electrochemistry The Basics of Organic Chemistry And more

Optimizing STEM Education With Advanced ICTs and Simulations Levin, Ilya, Tsybulsky, Dina, 2017-06-05 The role of technology in educational settings has become increasingly prominent in recent years When utilized effectively these tools provide a higher quality of learning for students Optimizing STEM Education With Advanced ICTs and Simulations is an innovative reference source for the latest scholarly research on the integration of

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Comprehension of Stoichiometric Concepts Connie Lynn Bannick Kemner,2007 Pp/Chemistry BarCharts, Inc.,2008-06-18

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