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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

Bond	kJ/mol	kJ/mol
C-H	413	431

Use the bond energies in Table 3.1 and $\Delta H = -201 \text{ kJ/mol}$ to calculate the bond energy of the C-Cl bond in kJ/mol .

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.

$$\begin{aligned}
 & (2 \times \text{C-H}) + 1(\text{C-C}) \\
 & 4 \times 413 + 1(340) = 2150 \text{ kJ/mol}
 \end{aligned}$$
- Calculate the energy released when bonds form in the products.

$$\begin{aligned}
 & \text{Bond broken} - \text{Bond formed} = -99 \\
 & 2150 - x = -99 \\
 & x = 2249 \text{ kJ/mol}
 \end{aligned}$$
- Calculate the energy of C-Cl bond.

$$\begin{aligned}
 & 6 \times (\text{C-H}) = 2478 \\
 & \text{N-H} = 2249 \\
 & \text{C} = 2478 + 2249 = 4727 \text{ kJ/mol}
 \end{aligned}$$

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Annelies Wilder-Smith

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Table of Contents 0620 32 M J 12 Mark Scheme

1. Understanding the eBook 0620 32 M J 12 Mark Scheme
 - The Rise of Digital Reading 0620 32 M J 12 Mark Scheme
 - Advantages of eBooks Over Traditional Books
2. Identifying 0620 32 M J 12 Mark Scheme
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 0620 32 M J 12 Mark Scheme
 - User-Friendly Interface
4. Exploring eBook Recommendations from 0620 32 M J 12 Mark Scheme
 - Personalized Recommendations

- 0620 32 M J 12 Mark Scheme User Reviews and Ratings
- 0620 32 M J 12 Mark Scheme and Bestseller Lists
- 5. Accessing 0620 32 M J 12 Mark Scheme Free and Paid eBooks
 - 0620 32 M J 12 Mark Scheme Public Domain eBooks
 - 0620 32 M J 12 Mark Scheme eBook Subscription Services
 - 0620 32 M J 12 Mark Scheme Budget-Friendly Options
- 6. Navigating 0620 32 M J 12 Mark Scheme eBook Formats
 - ePub, PDF, MOBI, and More
 - 0620 32 M J 12 Mark Scheme Compatibility with Devices
 - 0620 32 M J 12 Mark Scheme Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 0620 32 M J 12 Mark Scheme
 - Highlighting and Note-Taking 0620 32 M J 12 Mark Scheme
 - Interactive Elements 0620 32 M J 12 Mark Scheme
- 8. Staying Engaged with 0620 32 M J 12 Mark Scheme
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 0620 32 M J 12 Mark Scheme
- 9. Balancing eBooks and Physical Books 0620 32 M J 12 Mark Scheme
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 0620 32 M J 12 Mark Scheme
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 0620 32 M J 12 Mark Scheme
 - Setting Reading Goals 0620 32 M J 12 Mark Scheme
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 0620 32 M J 12 Mark Scheme
 - Fact-Checking eBook Content of 0620 32 M J 12 Mark Scheme

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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