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IGCSE A* MODEL ANSWER

0620 CHEMISTRY

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Paper 42



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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

Bond	N-H	N-O
Bond energy in kJ/mol	400	430

Use the bond energies in Table 3.1 and $\Delta H = -2140 \text{ kJ/mol}$ to calculate the bond energy of an N-O bond in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.

$$(400 \times 12) + 1(4 \times 400)$$

$$4800 + 1(4 \times 400)$$

$$3200 \text{ kJ}$$
- Calculate the energy released when bonds form in the products.

$$\Delta H_{\text{break}} - \Delta H_{\text{form}} = -90$$

$$3200 - x = -90$$

$$x = 3200 + 90$$

$$3290 \text{ kJ}$$
- Calculate the energy of an N-O bond.

$$6 \times (\text{N-O}) = 3290$$

$$\text{N-O} = \frac{3290}{6}$$

$$548 \text{ kJ/mol}$$

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L Darling-Hammond



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