Calculate the amount of heat needed to increase the temperature of 250 g of water from 20 °C to 46 °C.

q = m x C x ΔT q = 250 g x 4.18 J/g °C x 26°C q = 27,170 J or 27 kJ

Calculate the specific heat capacity of copper given that 204.75 J of energy raises the temperature of 15 g of copper from 25 °C to 60. °C.

q = m x C x ΔT C= q/m x ΔT C = 204.75 J / (15 g x 35 °C) C= 0.39 J/g °C

216 J of energy is required to raise the temperature of aluminum from 15 °C to 35 °C.
 Calculate the mass of aluminum. (Specific heat capacity of aluminum is 0.90 J/g °C).

q = m x C x ΔT m= q/C x ΔT m= 216 J/ (0.90 J/ g °C x 20 °C) m= 12 g

4. The initial temperature of 150 g of ethanol was 22 °C. What will be the final temperature of the ethanol if 3240 J was needed to raise the temperature of the ethanol? (Specific heat capacity of ethanol is 2.44 J/ g °C).

q = m x C x ΔT ΔT = q/m x C ΔT = 3240 J/ (150 g x 2.44 J/ g °C) ΔT = 8.85 °C Trust = 22 °C + 8.85 °C = 31 °C

5. When 435 J of heat is added to 3.4 g of olive oil at $21\,^{\circ}$ C, the temperature increases to 85 °C. What is the specific heat of the olive oil?

q = m x C x \(\Delta T\)
C= q/m x \(\Delta T\)
C = 435 \(\J/\) (3.4 g x 64 °C)
C = 2.0 \(\J/\) g °C

6. How many kilojoules of heat are absorbed when 1.00 L of water is heated from 18 °C to 85 °C? (Hint: You first need to determine the mass of the water, then calculate q in the requested unit.)

1 Specific Heat Answers

PT Brinkman

1 Specific Heat Answers:

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Table of Contents 1 Specific Heat Answers

- 1. Understanding the eBook 1 Specific Heat Answers
 - The Rise of Digital Reading 1 Specific Heat Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying 1 Specific Heat Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 1 Specific Heat Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from 1 Specific Heat Answers
 - Personalized Recommendations
 - 1 Specific Heat Answers User Reviews and Ratings
 - 1 Specific Heat Answers and Bestseller Lists

- 5. Accessing 1 Specific Heat Answers Free and Paid eBooks
 - 1 Specific Heat Answers Public Domain eBooks
 - 1 Specific Heat Answers eBook Subscription Services
 - 1 Specific Heat Answers Budget-Friendly Options
- 6. Navigating 1 Specific Heat Answers eBook Formats
 - o ePub, PDF, MOBI, and More
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 - 1 Specific Heat Answers Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of 1 Specific Heat Answers
 - Highlighting and Note-Taking 1 Specific Heat Answers
 - Interactive Elements 1 Specific Heat Answers
- 8. Staying Engaged with 1 Specific Heat Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 1 Specific Heat Answers
- 9. Balancing eBooks and Physical Books 1 Specific Heat Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 1 Specific Heat Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 1 Specific Heat Answers
 - Setting Reading Goals 1 Specific Heat Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 1 Specific Heat Answers
 - Fact-Checking eBook Content of 1 Specific Heat Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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