

**\*X10320\***

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| <b>Question Paper Code : X 10320</b> |
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*B.E./B.Tech. DEGREE EXAMINATIONS, NOVEMBER/DECEMBER 2020 AND  
APRIL/MAY 2021*

*Fifth/Sixth Semester*

*Computer Science and Engineering CS*

*8591 – COMPUTER NETWORKS*

*(Common to : Computer and Communication Engineering/Information  
Technology) (Regulations 2017)*

*Time : Three Hours*

*Maximum : 100 Marks*

*Answer ALL questions*

**PART – A**

**(10×2=20 Marks)**

1. *How are the subgroups of OSI model layers segregated by their functions ?*
2. *Differentiate Circuit-switched networks and Packet-switched networks.*
3. *List out the functions of the Data Link Layer.*
4. *Find the Hamming distance between two pair of code words : A*  
*= 01011*  
*B = 11110*
5. *Why IPv6 is preferred over IPv4 ?*
6. *Find the class of each address.*  
*a) 11000001 10000011 00011011 11111111*  
*b) 252.5.15.111*
7. *How congestion occurs in a network ?*
8. *What happens in the three way handshaking between any two devices ?*
9. *Define anonymous FTP.*
10. *What is the difference between IMAP and POP ?*

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**Chao Zhang**



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web the  $\text{Ag}^+ \text{aq}$  and  $\text{Cl}^- \text{aq}$  ions become  $\text{AgCl} \text{s}$  but the  $\text{Na}^+ \text{aq}$  ions and the  $\text{NO}_3^- \text{aq}$  ions stay as  $\text{Na}^+ \text{aq}$  ions and  $\text{NO}_3^- \text{aq}$  ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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web aug 8 2022 a spectator ion is an ion that does not take part in the chemical reaction and is found in solution both before and after the reaction the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction

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web the new format requires students to write balanced chemical equations showing only the reacting substances excluding spectator ions for three different sets of reactants and to answer a short question requiring no calculator about each reaction the questions are intended to test the students understanding of the meaning of the reactions

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web jun 2 2023 the spectator ions are cut out from a reaction hence the net ionic equation for the above reaction would be  $2\text{Cl}^- \text{aq} + \text{Cu}^{2+} \text{aq} \rightarrow \text{Cu} \text{s} + 2\text{Cl}^- \text{aq}$  these spectator ions are not actively involved in the reaction the existence of spectator ions affects the Debye length