

Arrhenius vs Bronsted-Lowry Acids and Bases

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Arrhenius

– acid = some substance that produces H^+ ions in an aqueous solution

– base = some substance that produces OH^- ions in an aqueous solution

Bronsted-Lowry

– acid = proton donor that gives up H^+

– base = proton acceptor that grabs the H^+

– conjugate acid = base + proton

– conjugate base = acid – proton

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Arrhenius Acids And Bronsted Lowry Acids Answers

NARAYAN CHANGDER



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