AP Chemistry Chapter 4 Practice Exam

- The Claus reactions, shown below, are used to generate elemental sulfur from hydrogen sulfide.

$$2H_2S + 3O_2 \rightarrow 2SO_2 + 2H_2O$$

 $SO_2 + 2H_2S \rightarrow 3S + 2H_2O$

How much sulfur (in grams) is produced from 21.5 grams of O₂? 21.5502 |mol 02 | 2503 | 35 | 32.0675 = 43.155

- A) 97.0 g
- B) 14.4 g
- C) 21.5 g
- none of these
- B
- Iron is produced from its ore by the reactions:

$$2C(e) + O_2(e) \rightarrow 2CO(e)$$

$$Fe_2O_3(s) + 3CO(g) \rightarrow 2Fe(s) + 3CO_3(g)$$

Closes:

How many moles of $O_2(g)$ are needed to produce 4.6 moles of Fe(s)?

- A) 2.3 mole O₂ B) 3.5 mole O-
- C) 4.6 mole O₂
- D) 6.9 mole O₂
- E) 13.8 mole O₂

- B
- Suppose the reaction Ca₂(PO₄)₂ + 3H₂SO₄ → 3CaSO₄ + 2H₂PO₄ is carried out starting with 153 g. of $Ca_2(PO_k)_2$

and 76.8 g of H_2SO_4 . How much phosphoric acid will be produced?

A) 76.7 gB) 51.1 g1 $C_{0.3}(C_{0.3}) = C_{0.3}(C_{0.3}) = C_{0.3}(C$

- B) 51.1 g C) 229.8 g

 - D) 115.1 c E) 96.7 g
- 7 6.89 H2504 | Ind H2504 2 H2 POY 1 3 =

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- - SO₂ reacts with H₂S as follows:

$$2H_2S + SO_2 \rightarrow 3S + 2H_2O$$

When 7.50 g of H₂S reacts with 12.75 g of SO₂, which statement applies?

- A) 6.38 g of sulfur are formed.
- B) 10.6 g of sulfur are formed.
- C) 0.0216 moles of H₂S remain.
- D) 1.13 g of H₂S remain.
- E) SO₂ is the limiting reagent.

4.6 mil Fersi 300 102 = 3.45mol02

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education It is hoped that this book will prove helpful to all involved in sharing the exciting story of the way humans have come to understand the molecular world one of the great examples of great human endeavour

The Science Teacher, 2009

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