

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive an empirical formula from a chemical structure without performing stoichiometric calculations.
2. To calculate the number of moles of a solid from a weighing without performing stoichiometric calculations.
3. To determine the molar mass of a solid.
4. To determine the stoichiometric amount of a solid reagent that would be produced.
5. To compare the theoretical amount to the actual amount of a solid reagent and calculate the percent yield.

Reagents and Apparatus:

Reagent of Sodium Chloride solution: 0.10 mol

Reagent of Calcium Chloride solution: 0.10 mol

Reagent of Silver Nitrate: 0.10 g

Reagent of Silver Nitrate: 0.10 g

Reagent of Silver Nitrate: 0.10 g

Reagent of Silver Nitrate: 0.10 g

Procedure:

Reagents and Calculations:

1. Calculations:

$$\text{moles of sodium chloride} = 0.10 \text{ mol NaCl} \times \frac{1.0}{100.0} = 0.0010 \text{ mol NaCl}$$

$$\text{moles of calcium chloride} = 0.10 \text{ mol CaCl}_2 \times \frac{1.0}{100.0} = 0.0010 \text{ mol CaCl}_2$$

$$\text{moles of silver nitrate produced} = 0.0010 \text{ mol} \times 2 = 0.0020 \text{ mol AgNO}_3$$

$$\text{mass of silver nitrate produced} = 0.0020 \text{ mol AgNO}_3 \times \frac{169.87 \text{ g}}{1 \text{ mol}} = 0.34 \text{ g AgNO}_3$$

2. Calculations: $\text{Na}_2\text{CO}_3 + \text{CaCl}_2 \rightarrow \text{CaCO}_3 + 2\text{NaCl}$

3. Calculations:

$$\text{mass of CaCO}_3 = 0.0010 \text{ mol} \times \frac{100.09 \text{ g}}{1 \text{ mol}} = 0.10 \text{ g CaCO}_3$$

$$\text{mass of CaCO}_3 = 0.10 \text{ g}$$

$$\text{mass of CaCO}_3 = 0.10 \text{ g} \times \frac{100.09 \text{ g}}{100.09 \text{ g}} = 0.10 \text{ g CaCO}_3$$

Answers For Lab 6b Heath Chemistry

S Marginson



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