

Name: Ahmed Faizan Date: 12 May Period:

Specific Heat and Heat Capacity Worksheet

DIRECTIONS: Use $q = (m)(C_p)(\Delta T)$ to solve the following problems. Show all work and units.

Ex: How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C?

$$\begin{aligned} q &= ? & \Delta t &= 55 - 22 = 33^\circ\text{C} & q &= m s \Delta t \\ m &= 10\text{g} & s &= 0.9 \text{ J/g}^\circ\text{C} & &= 10 \times 0.9 \times 33 \\ & & & & &= \underline{297 \text{ J}} \end{aligned}$$

1. The temperature of 335 g of water changed from 24.5°C to 26.4°C. How much heat did this sample absorb? C for water = 4.18 J/g°C

$$\begin{aligned} \Delta t &= 26.4 - 24.5 = 2 & q &= m s \Delta t \\ s &= 4.18 \text{ J/g}^\circ\text{C} & &= 335 \times 4.18 \times 2 \\ m &= 335\text{g} & &= \underline{2800.6 \text{ J}} \end{aligned}$$

2. How much heat in kilojoules has to be removed from 225g of water to lower its temperature from 25.0°C to 10.0°C?

$$\begin{aligned} \Delta t &= 25 - 10 = 15 & q &= m s \Delta t \\ m &= 225\text{g} & &= 225 \times 4.18 \times 15 \\ s &= 4.18 \text{ J/g}^\circ\text{C} & &= \underline{14107.5 \text{ J}} \\ 1\text{J} &= 1 \times 10^{-3} \text{ KJ} & \therefore & \underline{14.1075 \text{ KJ}} \end{aligned}$$

3. To bring 1.0kg of water from 25°C to 99°C takes how much heat input?

$$\begin{aligned} m &= 1\text{kg} = 1000\text{g} & q &= 1000 \times 4.18 \times 74 \\ \Delta t &= 99 - 25 = 74 & q &= \underline{309320 \text{ J}} \\ s &= 4.18 \text{ J/g}^\circ\text{C} & & \end{aligned}$$

An insulated cup contains 75.0g of water at 24.00°C. A 26.00g sample of metal at 82.25°C is added. The final temperature of the water and metal is 28.34°C. What is the specific heat of the metal?

Imp

$$\begin{aligned} (q_{\text{water}} &= -q_{\text{metal}}) & (C = m s) & \rightarrow s = \frac{C}{m} \\ q_{\text{water}} &= 75 \times 4.18 \times (28.34 - 24) & q &= m s \Delta t \\ &= 1252 & 1252 &= 26 \times s \times (82.25 - 28.34) \\ & & & 1252 = 26 \times 53.91 \times s \\ & & & s = \underline{0.893 \text{ J/g}^\circ\text{C}} \end{aligned}$$

5. A calorimeter has a heat capacity of 1265 J/°C. A reaction causes the temperature of the calorimeter to change from 22.34°C to 25.12°C. How many joules of heat were released in this process?

$$\begin{aligned} C &= 1265 \text{ J/}^\circ\text{C} \\ \Delta t &= 25.12 - 22.34 = 2.78 \\ q &= ? \\ q &= C \Delta t \\ &= 1265 \times 2.78 \\ &= \underline{3516.7 \text{ J}} \end{aligned}$$

Answers Of Specific Heat

William Abbott Stone



Answers Of Specific Heat:

Answers to Questions Set at Examinations Great Britain. Education Department. Department of Science and Art, 1889

Answers to questions set at the science and art examinations Science and art department, 1887 Matriculation model

answers: heat and light, magnetism and electricity. Being the papers from June 1888 to Jan. 1891, with answers by H.E.

Schmitz and R.W. Stewart London univ, exam. papers, 1891 *Fluid and Thermal Dynamics Answer Bank for Engineers*

Ethirajan Rathakrishnan, 2023-03-27 This book provides the essence of aerodynamics fluid mechanics experimental methods gas dynamics high enthalpy gas dynamics helicopter aerodynamics heat transfer and thermodynamics describing the underlying principles of these subjects before listing the set of multiple choice questions of each subject which will prove to be useful for engineering students to comfortably face and win in the competitive examinations for engineering studies engineering services civil services doctoral Degree program entrance and so on This book will also be of value for those facing job interviews for academic positions in universities and research organizations or laboratories **Numerical**

Physics With Chapterwise Question - Answers Class XI by D. C. Upadhyay, Dr. J. P. Goel, Er. Meera Goyal D. C.

Upadhyay, Dr. J. P. Goel, Er. Meera Goyal, 2020-06-27 Unit 1 Physical World and Measurement 1 Systems of Units and Measurement 2 Significant Figures and Error Analysis 3 Dimensional Analysis Unit 2 Kinematics 4 Motion in a Straight Line 5 Vector Analysis 6 Motion in a Plane Unit 3 Laws of Motion 7 Newton's Laws of Motion 8 Friction 9 Uniform Circular Motion Unit 4 Work Energy and Power 10 Work Energy and Power Unit 5 Motion of System of Particles and Rigid Body 11 Centre of Mass 12 Rotational Motion and Moment of Inertia Unit 6 Gravitation 13 Gravitation Unit 7 Properties of Bulk Matter 14 Elasticity 15 Pressure of Fluids 16 Viscosity 17 Surface Tension 18 Temperature and Calorimetry 19 Transfer of Heat Unit 8 Thermodynamics 20 First Law of Thermodynamics 21 Second Law of Thermodynamics Unit 9 Behaviour of Perfect Gases and Kinetic Theory of Gases 22 Behaviour of Perfect Gas and Kinetic Theory of Gases Unit 10 Oscillations and Waves 23 Oscillations 24 Speed of Mechanical Waves and Progressive Waves 25 Superposition of Waves Interference and Beats 26 Reflection of Waves Stationary Waves in Stretched Strings and Organ Pipes 27 Doppler's Effect *Physics with Answers*

Andrew R. King, Oded Regev, 1997-05-28 This book contains 500 problems covering all of introductory physics along with clear step by step solutions to each problem **Numerical Examples in Heat** Richard Evan Day, 1885 *Experimental*

physics William Abbott Stone, 1899 **Review Guide for LPN/LVN Pre-entrance Exam** National League for Nursing, Natl League Nursing, 2008-09-12 Providing an overview of the math science and reading comprehension skills necessary for

admission to LPN programs this bestselling guide includes multiple choice questions and practice examinations in each of the three test areas *Physics Question Bank IIT JEE Mains - Chapterwise*, Physics Question Bank IIT JEE Mains Table of

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Lessons in Elementary Physics Balfour Stewart, 1895

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web oct 10 2009 was playing a fresh game with some newer players and there was a small teeny tiny chance i threw too many goblins at them infinitesimal chance really once i realized what was happening after nearly killing half the party i fudged the hp on the remaining ones and once about half of them were dead the rest retreated

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