

5. Gator Electronics manufactures a critical electronic component used in high definition television sets using labor and state-of-the-art high-tech equipment. Each manufactured component is classified either "A-Quality" or "B-Quality" or "C-Quality." An "A-Quality" component is sold for \$120; a "B-Quality" component is sold for \$70 and "C-quality" components are scrapped because they have no market value. In 2016, Gator Electronics has launched a Total Quality Program to improve the quality of its product and to attain a better competitive standing. Production data for 2016 and 2017 are summarized below.

	<u>2016</u>	<u>2017</u>
Total Number of Components Produced	1100	1400
A-Quality	900	1200
B-Quality	150	180
C-Quality	50	20
Labor hours	3000	3480
Machine hours	2000	2610
Labor Cost	\$10/hour	\$10/hour
Machine Cost	\$24.50/hour	\$24.50/hour

Gator Electronics uses the following productivity measures:

Labor Productivity = \$ Sales/Labor hour

Machine Productivity = \$ Sales/Machine hour

Total Productivity = \$ Sales/\$ of input

Compute the labor, machine and total productivity for 2016 and 2017 and interpret your results.

Answer To Deloitte Gator Electronics

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web calculate the mass percent mole fraction molality and molarity of hexane C_6H_{14} d 0 660 g cm^3 when it is mixed with heptane C_7H_{16} d 0 684 g cm^3 in a 35 0 55 0 ml ratio assume that the volumes are additive

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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water 2 what is the mole fraction of oxygen in a gaseous mixture containing 25 grams of oxygen 15 grams of

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web a derivatization that achieves this goes as molarity moles of solute volume of solution in liter moles molarity x volume in liter mass moles x molar mass numerical answers 1

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web what is the mole fraction of licl in the solution 11 8 practice problem acetone c 3 h 6 o has a density of 0 791 g ml and xylene c 8 h 10 has a density of 0 864 g ml a solution was prepared by mixing acetone and xylene calculate the mole fraction of acetone in a solution if 19 0 ml of acetone was dissolved in 98 4 ml of xylene 7 1

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web since the mole fraction of the o 2 is 0 050 it contributes 0 050 or if you will 5 0 of the total pressure thus total pressure is 0 21 0 050 4 2 atm problem 9 a sample of 1 43 g of helium and an unweighed quantity of o 2 are mixed in a flask at room temperature