অধ্যায় ১: ল্যাবরেটরির নিরাপদ ব্যবহার

চমক! পরীক্ষাটি কঠিন, এখানে কোর্স পরীক্ষা, ভর্তিতে ফলোয়ারদের শীর্ষস্থানীয় কলেজসমূহের পরীক্ষা এবং মডেল টেস্টের প্রশ্নগুলো পুরোপুরি উচ্চতর অধ্যায়ভিত্তিক দেওয়া হয়েছে। এগুলো অনুশীলন করলে তুমি এ অধ্যায়টি থেকে যেকোনো সৃজনশীল প্রশ্নসমূহের এখনই উত্তর লিখতে পারবে সন্দেহই।

প্রশ্ন ▶ ১

পাত্র-A

পাত্র-B

নীচে দেখ ১১ / এক্স বহু ১১

ক. মোলার দ্রবণ কী? ১

খ. "কোনাস পানিমুক্ত লবণের দ্রাব্যতা, তাপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়"— ব্যাখ্যা করো। ২

গ. A-পাত্রে বিদ্যমান এসিডের ভর কত নির্ণয় করো। ৩

ঘ. দ্রবণ দুটি মিশ্রিত করলে, মিশ্রণটি কী পরিমাণ অম্লীয় বা ক্ষারীয় হবে— বিশ্লেষণ করো। ৪

১ নং প্রশ্নের উত্তর

ক 1000 mL দ্রবণে কোন দ্রবের 1 mol দ্রবীভূত থাকলে উক্ত দ্রবণকে মোলার দ্রবণ বলে।

খ কোনাস পানিমুক্ত লবণের দ্রাব্যতা তাপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়। এর কারণ, তাপমাত্রা বৃদ্ধির সাথে সাথে দ্রাব্যতা বৃদ্ধি পেতে থাকলেও একটি নির্দিষ্ট তাপমাত্রায় কোনাস পানি মুক্ত লবণ নিষ্কাশিত হয়। তখন দ্রাব্যতা হ্রাস পেতে থাকে। উদাহরণস্বরূপ $Na_2SO_4.10H_2O$ এর কথা বলা যায়। $Na_2SO_4.10H_2O$ এর ক্ষেত্রে পানিতে দ্রবীভূত অবস্থার সময় বিয়োজন প্রক্রিয়াটি তাপশ্রয়ী প্রক্রিয়া কিন্তু 32.4°C তাপমাত্রার উপরে $Na_2SO_4.10H_2O$ যখন Na_2SO_4 এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি তাপোৎপাদী হয়। এ কারণে 32.4°C তাপমাত্রার উপরে তাপমাত্রা বৃদ্ধিজনিত কারণে দ্রাব্যতার হ্রাস ঘটে।

গ

আমরা জানি,
$w = SMV$
$= (0.5 \times 98 \times 0.25)$ g
$= 12.25$ g
অতএব, দ্রবীভূত H_2SO_4 এর ভর $= 12.25$ g

উদ্দীপকের A-পাত্রে অবস্থিত,
H_2SO_4 এর আণবিক ভর,
$M = 1 \times 2 + 32 + 16 \times 4$
$= 2 + 32 + 64$
$= 98$ g/mol
দ্রবণের ঘনমাত্রা,
$S = 0.5$ M
দ্রবণের আয়তন,
$V = 250$ mL
$= \frac{250}{1000} = 0.25$L
দ্রবীভূত H_2SO_4 এর ভর, $w = ?$

ঘ পাত্র-A: H_2SO_4 এর মোল সংখ্যা $= \left(0.5 \times \frac{250}{1000}\right)$ mol
$= 0.125$ mol
পাত্র-B: Na_2CO_3 এর মোল সংখ্যা $= \left(0.11 \times \frac{200}{1000}\right)$ mol
$= 0.077$ mol
সমীকৃত বিক্রিয়া: $H_2SO_4 + Na_2CO_3 \rightarrow Na_2SO_4 + H_2O + CO_2$
সমীকরণ হতে পাই, 1 mol H_2SO_4 = 1 mol Na_2CO_3
1 mol Na_2CO_3 বিক্রিয়া করে H_2SO_4 এর 1mol এর সাথে
0.077 mol " " " " (0.077) mol " "
" 0.077 mol " "
অবশিষ্ট H_2SO_4 এর পরিমাণ $= (0.125 - 0.077)$ mol
$= 0.048$ mol
পাত্র-A এর দ্রবণের সাথে পাত্র-B এর দ্রবণের বিক্রিয়ার ফলে 0.0 mol H_2SO_4 অবশিষ্ট থাকায় মিশ্রিত দ্রবণের প্রকৃতি হবে অম্লীয়।

প্রশ্ন ▶ ২ একজন ছাত্র ল্যাবরেটরিতে X-পদ্ধতিতে 10 mg এর চেয়ে ক পরিমাণ এবং Y-পদ্ধতিতে 0.5 g পরিমাণ 2% $\frac{W}{V}$ HCl এর দ্রবণ ত্ব করে।

নীচে দেখ ১১ / এক্স

ক. নোড কাকে বলে?

খ. Cl_2 অপোলার— ব্যাখ্যা করো।

গ. উদ্দীপকের এসিডটির মোলার ঘনমাত্রা নির্ণয় করো।

ঘ. উদ্দীপকের কোন পদ্ধতি পরিবেশবান্ধব? বিশ্লেষণ করো।

২ নং প্রশ্নের উত্তর

ক নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেকট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে নোড বলে।

খ সমযোজী বন্ধনে আবদ্ধ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা সমান হলে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হলে সমযোজী বন্ধন বিশুদ্ধ ও অপোলার হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর কম হলে অপোলার হয়। যেহেতু Cl এর তড়িৎ ঋণাত্মকতা 3.0 ও Cl অণুতে তড়িৎ ঋণাত্মকতার পার্থক্য $(3 - 3) = 0$, তাই Cl_2 বিশুদ্ধ সমযোজী ও অপোলার যৌগ। উল্লেখ্য যৌগিক অণুর ক্ষেত্রে এরূপ হয় থাকে।

গ উদ্দীপকে বর্ণিত এসিডটি হলো 2% $\left(\frac{W}{V}\right)$ HCl।
2% $\left(\frac{W}{V}\right)$ HCl এর অর্থ হলো—
100 mL দ্রবণে HCl এর ভর 2g
∴ 1000 mL " " " " $\left(\frac{2 \times 1000}{100}\right)$ g
$= 20$ g

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