

Stoichiometry Worksheet #2 (mole-mass, mass-mole problems)



- a. If 15.0g of N_2O_4 was produced, how many moles of O_2 were required?

$$\frac{15.0\text{g N}_2\text{O}_4}{92.0\text{g N}_2\text{O}_4} \left| \frac{1 \text{ mol N}_2\text{O}_4}{1 \text{ mol N}_2\text{O}_4} \right| \left| \frac{2 \text{ mol O}_2}{1 \text{ mol N}_2\text{O}_4} \right| = 0.326 \text{ mol O}_2$$

- b. If 4.0×10^{-3} moles of oxygen reacted, how many grams of N_2 were needed?

$$\frac{4.0 \times 10^{-3} \text{ mol O}_2}{2 \text{ mol O}_2} \left| \frac{1 \text{ mol N}_2}{1 \text{ mol O}_2} \right| \left| \frac{28.0\text{g N}_2}{1 \text{ mol N}_2} \right| = 5.6 \times 10^{-2} \text{ g N}_2$$



- a. What is the mass of potassium nitrate that is produced when 2.04 moles of potassium phosphate react?

$$\frac{2.04 \text{ mol K}_3\text{PO}_4}{1 \text{ mol K}_3\text{PO}_4} \left| \frac{3 \text{ mol KNO}_3}{1 \text{ mol K}_3\text{PO}_4} \right| \left| \frac{101.1 \text{ g KNO}_3}{1 \text{ mol KNO}_3} \right| = 619\text{g KNO}_3$$

- b. If 5.80g of aluminum phosphate are formed, how many moles of aluminum nitrate reacted?

$$\frac{5.80\text{g AlPO}_4}{122.0\text{g AlPO}_4} \left| \frac{1 \text{ mol AlPO}_4}{1 \text{ mol AlPO}_4} \right| \left| \frac{1 \text{ mol Al}(\text{NO}_3)_3}{1 \text{ mol AlPO}_4} \right| = 4.75 \times 10^{-2} \text{ mol Al}(\text{NO}_3)_3$$



- a. If you have 5.50 mol of CaC_2 , how much C_2H_2 do you get?

$$\frac{5.50 \text{ mol CaC}_2}{1 \text{ mol CaC}_2} \left| \frac{1 \text{ mol C}_2\text{H}_2}{1 \text{ mol CaC}_2} \right| \left| \frac{26.0\text{g C}_2\text{H}_2}{1 \text{ mol C}_2\text{H}_2} \right| = 143\text{g C}_2\text{H}_2$$

- b. How many moles of water are needed when 65.0g of CaC_2 have reacted?

$$\frac{65.0\text{g CaC}_2}{64.1\text{g CaC}_2} \left| \frac{1 \text{ mol CaC}_2}{1 \text{ mol CaC}_2} \right| \left| \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol CaC}_2} \right| = 2.03 \text{ mol H}_2\text{O}$$

4. In photosynthesis, water reacts with carbon dioxide to give oxygen and glucose ($\text{C}_6\text{H}_{12}\text{O}_6$). Write and balance the chemical equation. How many moles of CO_2 are required to make 120.0g of glucose?



$$\frac{120.0\text{g C}_6\text{H}_{12}\text{O}_6}{180.0 \text{ g C}_6\text{H}_{12}\text{O}_6} \left| \frac{1 \text{ mol C}_6\text{H}_{12}\text{O}_6}{1 \text{ mol C}_6\text{H}_{12}\text{O}_6} \right| \left| \frac{6 \text{ mol CO}_2}{1 \text{ mol C}_6\text{H}_{12}\text{O}_6} \right| = 4.000 \text{ mol CO}_2$$

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Instructor's Guide for Introductory Chemistry in the Laboratory James F. Hall,1996 A Concrete Stoichiometry Unit for High School Chemistry Jennifer Louise Pakkala,2006 Improving Student Comprehension in Chemistry Laboratories Tracy Lynn Haroff,2006 Regents Exams and Answers: Chemistry--Physical Setting Revised Edition Barron's Educational Series,Albert Tarendash,2021-01-05 Barron's Regents Exams and Answers Chemistry provides essential practice for students taking the Chemistry Regents including actual recently administered exams and thorough answer explanations for all questions This book features Eight actual administered Regents Chemistry exams so students can get familiar with the test Thorough explanations for all answers Self analysis charts to help identify strengths and weaknesses Test taking techniques and strategies A detailed outline of all major topics tested on this exam A glossary of important terms to know for test day

CBSE Chemistry Chapterwise Case Study Class 11 Priti Singhal,2024-11-17 This book is structured to align with the latest syllabus and curriculum guidelines ensuring that the content is both relevant and rigorous Each chapter begins with a clear set of learning objectives providing a roadmap for students to understand what they will achieve by the end of the chapter We have included numerous diagrams illustrations and real life examples to make complex concepts more accessible and engaging

Whole-class Inquiry Dennis W. Smithenry,Joan Gallagher-Bolos,2009 In response to requests from science education professionals this is the perfect vehicle for implementing and assessing this concept of whole class inquiry in your classroom This is a must have package for preservice and inservice middle and high school science teachers Barron's Science 360: A Complete Study Guide to Chemistry with Online

Practice Barron's Educational Series, Mark Kernion, Joseph A. Mascetta, 2021-09-07 provides a complete guide to the fundamentals of chemistry Page 4 of cover **Multiple Solution Methods for Teaching Science in the Classroom** Stephen DeMeo, 2008 For the first time in science education the subject of multiple solution methods is explored in book form While a multiple method teaching approach is utilized extensively in math education there are very few journal articles and no texts written on this topic in science Teaching multiple methods to science students in order to solve quantitative word problems is important for two reasons First it challenges the practice by teachers that one specific method should be used when solving problems Secondly it calls into question the belief that multiple methods would confuse students and retard their learning Using a case study approach and informed by research conducted by the author this book claims that providing students with a choice of methods as well as requiring additional methods as a way to validate results can be beneficial to student learning A close reading of the literature reveals that time spent on elucidating concepts rather than on algorithmic methodologies is a critical issue when trying to have students solve problems with understanding It is argued that conceptual understanding can be enhanced through the use of multiple methods in an environment where students can compare evaluate and verbally discuss competing methodologies through the facilitation of the instructor This book focuses on two very useful methods proportional reasoning PR and dimensional analysis DA These two methods are important because they can be used to solve a large number of problems in all of the four academic sciences biology chemistry physics and earth science This book concludes with a plan to integrate DA and PR into the academic science curriculum starting in late elementary school through to the introductory college level A challenge is presented to teachers as well as to textbook writers who rely on the single method paradigm to consider an alternative way to teach scientific problem solving

Cooperative Chemistry Lab Manual Cooper, 2005-02 The laboratory course described in the lab manual emphasizes experimental design data analysis and problem solving Inherent in the design is the emphasis on communication skills both written and oral Students work in groups on open ended projects in which they are given an initial scenario and then asked to investigate a problem There are no formalized instructions and students must plan and carry out their own investigations

Chemistry: The Easy Way Joseph A. Mascetta, Mark Kernion, 2019-08-06 A self teaching guide for students Chemistry The Easy Way provides easy to follow lessons with comprehensive review and practice This edition features a brand new design and new content structure with illustrations and practice questions An essential resource for High school and college courses Virtual learning Learning pods Homeschooling Chemistry The Easy Way covers Atomic Structure Chemical Formulas Electrochemistry The Basics of Organic Chemistry And more **Optimizing STEM Education With Advanced ICTs and Simulations** Levin, Ilya, Tsybulsky, Dina, 2017-06-05 The role of technology in educational settings has become increasingly prominent in recent years When utilized effectively these tools provide a higher quality of learning for students Optimizing STEM Education With Advanced ICTs and Simulations is an innovative reference source for the latest scholarly research on

the integration of digital tools for enhanced STEM based learning environments Highlighting a range of pivotal topics such as mobile games virtual labs and participatory simulations this publication is ideally designed for educators professionals academics and students seeking material on emerging educational technologies *Laboratory Experiments for General Chemistry* Harold R. Hunt,Toby F. Block,George M. McKelvy,2002 This established manual focuses on using non hazardous materials to teach the experimental nature of general chemistry Experiments are written to address students of various academic backgrounds and differing interests and abilities in chemistry While most experiments can be conducted in a single three hour period some have been designed to be completed over an extended time to illustrate that chemical systems do not work at an arbitrary schedule Suggestions are provided for combining experiments of shorter length and similar pedagogy

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color illustrations and photos Illustrated Guide to Home Chemistry Experiments offers introductory level sessions suitable for a middle school or first year high school chemistry laboratory course and more advanced sessions suitable for students who intend to take the College Board Advanced Placement AP Chemistry exam A student who completes all of the laboratories in this book will have done the equivalent of two full years of high school chemistry lab work or a first year college general chemistry laboratory course This hands on introduction to real chemistry using real equipment real chemicals and real quantitative experiments is ideal for the many thousands of young people and adults who want to experience the magic of chemistry *Energy Research Abstracts* ,1989

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