



Lab Data



Verify your calculation. In your set-up, do your units cancel?

Mass of magnesium (g)	.015
Moles of magnesium (mol)	.00062
Temperature of water (°C)	23.0
Temperature of water (K)	296.15
Vapor pressure of water (mmHg)	21.1
Barometric pressure (mmHg)	778.0

Observations

I observe that there are bubbles that are visible.

Volume of hydrogen gas collected (mL)	15.00
Volume of hydrogen gas collected (L)	0.01500
Pressure of hydrogen gas (mmHg)	756.9
Pressure of hydrogen gas (atm)	0.9959
Moles of hydrogen gas produced (mol)	.00690
Ideal gas constant (L · atm · mol ⁻¹ · K ⁻¹)	0.080

Answers Escience Lab Ideal Gas Law

Robert Boyle, Emile Hilaire Amagat



Answers Escience Lab Ideal Gas Law:

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Hilaire Amagat, 2023-07-18 This scientific memoir covers the fundamental laws and properties of gases as understood by three pioneering researchers in the field Through experiments and rigorous analysis the authors explore topics such as temperature pressure heat and chemical reactions shedding light on the complex and fascinating dynamics of gas molecules With detailed descriptions and illustrations this book provides a comprehensive introduction to the subject for students and scholars alike This work has been selected by scholars as being culturally important and is part of the knowledge base of civilization as we know it This work is in the public domain in the United States of America and possibly other nations Within the United States you may freely copy and distribute this work as no entity individual or corporate has a copyright on the body of the work Scholars believe and we concur that this work is important enough to be preserved reproduced and made generally available to the public We appreciate your support of the preservation process and thank you for being an important part of keeping this knowledge alive and relevant *Aberrations from the Ideal Gas Laws in Systems of One and Two Components* Otto Maass, 1926 **Aberrations from the Ideal Gas Laws and a Precision Method for the Determination of the Densities of Gases** Douglas Le Baron Peters Cooper, 1931

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