

AP Chemistry Chapter 4 Practice Exam

D

1. The Claus reactions, shown below, are used to generate elemental sulfur from hydrogen sulfide.

How much sulfur (in grams) is produced from 21.5 grams of O_2 ?

A) 97.0 g

B) 14.4 g

C) 21.5 g

D) 43.1 g

E) none of these

$$\frac{21.5 \text{ g O}_2}{32.00 \text{ g O}_2} \times \frac{1 \text{ mol O}_2}{1 \text{ mol O}_2} \times \frac{2 \text{ SO}_2}{3 \text{ O}_2} \times \frac{3 \text{ S}}{\text{SO}_2} \times \frac{32.06 \text{ g S}}{1 \text{ mol S}} = 43.1 \text{ g S}$$

B

2. Iron is produced from its ore by the reactions:

How many moles of $\text{O}_2(g)$ are needed to produce 4.6 moles of $\text{Fe}(s)$?A) 2.3 mole O_2 B) 3.5 mole O_2 C) 4.6 mole O_2 D) 6.9 mole O_2 E) 13.8 mole O_2

$$\frac{4.6 \text{ mol Fe}(s)}{2 \text{ Fe}} \times \frac{3 \text{ CO}}{2 \text{ Fe}} \times \frac{1 \text{ O}_2}{2 \text{ CO}} = 3.45 \text{ mol O}_2$$

B

3. Suppose the reaction
- $\text{Ca}_3(\text{PO}_4)_2 + 3\text{H}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{H}_3\text{PO}_4$
- is carried out starting with 153 g of
- $\text{Ca}_3(\text{PO}_4)_2$

and 76.8 g of H_2SO_4 . How much phosphoric acid will be produced?

A) 76.7 g

B) 31.1 g

C) 229.8 g

D) 115.1 g

E) 96.7 g

$$\frac{153 \text{ g Ca}_3(\text{PO}_4)_2}{310.18 \text{ g Ca}_3(\text{PO}_4)_2} \times \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} \times \frac{2 \text{ H}_3\text{PO}_4}{1 \text{ Ca}_3(\text{PO}_4)_2} = 0.97 \text{ mol H}_3\text{PO}_4$$

$$\frac{76.8 \text{ g H}_2\text{SO}_4}{98.08 \text{ g H}_2\text{SO}_4} \times \frac{1 \text{ mol H}_2\text{SO}_4}{1 \text{ mol H}_2\text{SO}_4} \times \frac{2 \text{ H}_3\text{PO}_4}{3 \text{ H}_2\text{SO}_4} = 0.51 \text{ mol H}_3\text{PO}_4$$

B

- 4.
- SO_2
- reacts with
- H_2S
- as follows:

When 7.50 g of H_2S reacts with 12.75 g of SO_2 , which statement applies?

A) 6.38 g of sulfur are formed.

B) 10.6 g of sulfur are formed.C) 0.0216 moles of H_2S remain.D) 1.13 g of H_2S remain.E) SO_2 is the limiting reagent.

$$\frac{7.50 \text{ g H}_2\text{S}}{34.03 \text{ g H}_2\text{S}} \times \frac{1 \text{ mol H}_2\text{S}}{1 \text{ mol H}_2\text{S}} \times \frac{3 \text{ S}}{2 \text{ H}_2\text{S}} \times \frac{32.06 \text{ g S}}{1 \text{ mol S}} = 10.89 \text{ g S}$$

$$\frac{12.75 \text{ g SO}_2}{64.06 \text{ g SO}_2} \times \frac{1 \text{ mol SO}_2}{1 \text{ mol SO}_2} \times \frac{3 \text{ S}}{1 \text{ SO}_2} \times \frac{32.06 \text{ g S}}{1 \text{ mol S}} = 19.13 \text{ g S}$$

Answers For Ap Chemistry Test Open Ended

Rosina Ehmann



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Children's Reading Comprehension and Assessment Scott G. Paris, Steven A. Stahl, 2005-03-23 The volume identifies critical issues and cutting edge research in assessing children's reading comprehension

Socioculturally Responsive Assessment Randy E. Bennett, Linda Darling-Hammond, Aneesha Badrinarayan, 2025-05-01 Socioculturally Responsive Assessment assembles the best available thinking from within and outside the educational measurement community about the theoretical foundations and systems level policy implications of formal assessment programs designed to be socioculturally responsive Synthesized from culturally responsive assessment design and practices culturally relevant pedagogy and funds of knowledge universal design for learning the learning sciences and other literatures this emerging concept affirms that students learning and performance is inextricably tied to the social cultural and linguistic contexts in which they live and develop knowledge Across four sections this book provides an argument and initial evidence for impact on students users and assessment quality offers guidance for implementation and examines the potential limitations pitfalls barriers and measurement issues that such programs will inevitably raise Scholars teaching faculty test developers and policymakers will come away with integral foundations new assessment approaches and a greater sense of the potential for positive impact that these assessments may afford

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Chemistry Steven S. Zumdahl,1993 **Visualization: Theory and Practice in Science Education** John K. Gilbert,Miriam Reiner,Mary Nakhleh,2007-12-05 External representations pictures diagrams graphs concrete models have always been valuable tools for the science teacher This book brings together the insights of practicing scientists science education researchers computer specialists and cognitive scientists to produce a coherent overview It links presentations about cognitive theory its implications for science curriculum design and for learning and teaching in classrooms and laboratories **Green Organic Chemistry and its Interdisciplinary Applications** Vera M. Kolb,2017-04-21 Green Organic Chemistry and Its Interdisciplinary Applications covers key developments in green chemistry and demonstrates to students that the developments were most often the result of innovative thinking Using a set of selected experiments all of which have been performed in the laboratory with undergraduate students it demonstrates how to optimize and develop green experiments The book dedicates each chapter to individual applications such as Engineering The chemical industry The pharmaceutical industry Analytical chemistry Environmental chemistry Each chapter also poses questions at the end with the answers included By focusing on both the interdisciplinary applications of green chemistry and the innovative thinking that has produced new developments in the field this book manages to present two key messages in a manner where they reinforce each other It provides a single and concise reference for chemists instructors and students for learning about green organic chemistry and its great and ever expanding number of applications **The American Drug Clerks Journal** ,1887 **Journal of Applied Chemistry** ,1869 *The Education Index* ,1990 American Druggists' Circular and Chemical Gazette ,1892 **Resources in Education** ,1994 **Cumulated Index Medicus** ,2000 A Complete Crash Course in AIEEE 2011 K.K. Arora, Dinesh Khattar, Ravi Raj Dudeja, **Industries** ,1887

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