

11.1

DESCRIBING CHEMICAL REACTIONS

Section Review

Objectives

- 1. Represent the processes of a reaction.
- 2. Explain how to write a balanced equation.
- 3. List the steps for writing a complete chemical equation.

Vocabulary

- 1. chemical equation
- 2. chemical equation
- 3. reactant
- 4. coefficient
- 5. balanced equation

Part A Completion

Write the coefficient for each in the balanced chemical equation. Label each product and reactant as a compound or element.

1. A chemical equation is always represented by a balanced _____ 1. _____
 The chemical formula for a chemical compound is _____ 2. _____ The mass
 number for an element is represented by the _____ 3. _____ An element has only
 one type of atom. It is _____ 4. _____ A chemical equation is a reaction that
 shows the change in the _____ 5. _____ of the reactants and products.
 The _____ 6. _____ of a chemical equation is the number of atoms of each element on
 each side of the equation. The equation is balanced if the number of atoms of each element is
 the same on both sides of the equation.

2. The chemical formula for a compound is always represented by a balanced _____ 1. _____
 The chemical formula for a compound is _____ 2. _____ The mass number for an
 element is represented by the _____ 3. _____ An element has only one type of
 atom. It is _____ 4. _____ A chemical equation is a reaction that shows the
 change in the _____ 5. _____ of the reactants and products.
 The _____ 6. _____ of a chemical equation is the number of atoms of each element on
 each side of the equation. The equation is balanced if the number of atoms of each element is
 the same on both sides of the equation.

- 1. _____
- 2. _____
- 3. _____
- 4. _____
- 5. _____
- 6. _____
- 7. _____
- 8. _____
- 9. _____
- 10. _____
- 11. _____
- 12. _____
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- 18. _____
- 19. _____
- 20. _____

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