

Nuclear Decay

Types of Radiation

1. Complete the table below to compare the properties of alpha, beta, and gamma radiation.

Name	Alpha	Beta	Gamma
Greek letter	α	β -	γ
Symbol	${}^4_2\text{He}$	${}^0_{-1}e$	γ
Composition	2 protons, 2 neutrons	an electron	radiation (light)
Charge	0	-1	0
Stopped by	paper, skin, clothing	aluminum	lead

Nuclear Decay Reactions

In problems 2 – 5, fill in the missing particle for each nuclear decay reaction. Then, match each equation to the appropriate type of nuclear decay.

- _____ 2. ${}^{14}_6\text{C} \rightarrow {}^{14}_7\text{N} + \text{_____}$ beta A. alpha emission
- _____ 4. ${}^{218}_{84}\text{Po} \rightarrow {}^{214}_{82}\text{Pb} + \text{_____}$ alpha B. beta emission
- _____ 5. $\text{_____} + {}^0_{-1}e \rightarrow {}^{147}_{60}\text{Nd}$ electron capture C. electron capture

In problems 6 and 7, write a balanced equation for each nuclear decay reaction.

6. Decay of radium-223 by alpha (α) emission.



7. Decay of iodine-126 by electron (β) emission.



8. How are the atomic number and mass number of a nuclide affected by Beta Decay?

The mass number does not change but the atomic number increases by 1.

9. How are the atomic number and mass number of a nuclide affected by the emission of gamma radiation?

The atomic number and mass number do not change.

10. Which has a slightly greater mass in a nuclear reaction, the reactants or the products? Why?

It depends on the type of reaction. A fusion reaction will have greater mass in the products because it is combining two nuclei. A fission reaction will not change mass.

2radioactive Decay Worksheet Answer Key

Michael J. Padilla

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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1. Understanding the eBook 2radioactive Decay Worksheet Answer Key
 - The Rise of Digital Reading 2radioactive Decay Worksheet Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying 2radioactive Decay Worksheet Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 2radioactive Decay Worksheet Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from 2radioactive Decay Worksheet Answer Key
 - Personalized Recommendations
 - 2radioactive Decay Worksheet Answer Key User Reviews and Ratings
 - 2radioactive Decay Worksheet Answer Key and Bestseller Lists
5. Accessing 2radioactive Decay Worksheet Answer Key Free and Paid eBooks
 - 2radioactive Decay Worksheet Answer Key Public Domain eBooks
 - 2radioactive Decay Worksheet Answer Key eBook Subscription Services
 - 2radioactive Decay Worksheet Answer Key Budget-Friendly Options
6. Navigating 2radioactive Decay Worksheet Answer Key eBook Formats

- ePub, PDF, MOBI, and More
 - 2radioactive Decay Worksheet Answer Key Compatibility with Devices
 - 2radioactive Decay Worksheet Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 2radioactive Decay Worksheet Answer Key
 - Highlighting and Note-Taking 2radioactive Decay Worksheet Answer Key
 - Interactive Elements 2radioactive Decay Worksheet Answer Key
 8. Staying Engaged with 2radioactive Decay Worksheet Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 2radioactive Decay Worksheet Answer Key
 9. Balancing eBooks and Physical Books 2radioactive Decay Worksheet Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 2radioactive Decay Worksheet Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine 2radioactive Decay Worksheet Answer Key
 - Setting Reading Goals 2radioactive Decay Worksheet Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of 2radioactive Decay Worksheet Answer Key
 - Fact-Checking eBook Content of 2radioactive Decay Worksheet Answer Key
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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 - Interactive and Gamified eBooks

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