External Semester Examination Fall, 2022-23 BS Chemistry (5th Semester)

Objective

CHM 505 Time: 30min Total Marks: (20)

Physical Chemistry I Q No 01. Choose the correct option

The temperature at which the second coefficient of Virial equation become zero is known as

a. Critical temperature b. Boyle's temperature 2. Average distance of gas molecules between two successive collisions c. Eutectic point d. Boiling point

a. Mean path b. path length c. free path

3. Vrms for O₂ is -----times slower than that of H₂ a. 4 times. b. 8 times c. 16 times

d. 32 times At high temperatures the number of molecules possessing most probable velocities

a. Increase b. decrease c. not affected d. none of above

5. Based on the first law of thermodynamics, which one of the following is correct? a. For an isothermal process, q=+w

b. For an isochoric process, q= -AE For an adiabatic process, w= -ΔΕ d. For a cyclic process, q= -w

6. One mole of which of the following has the highest entropy?

a. liquid Nitrogen b. Hydrogen gas c. Meroury 7. An engine operating between 100 and OC, takes 453.6kcal of heat, useful work done is

a.453kJ b. 463kJ c. 509kJ d. 408k3

8. Which of the following is not an intensive property

a. Temperature b. surface tension viscosity. d. entropy

9. For a spontaneous process.

a. AG-0 b. AGH -ve c. AG=+ve d. 45~0

10. An isolated system comprises the liquid in equilibrium with vapor phase, Molar entropy of vapor is

a. More than that of liquid b. Less than that of liquid c. Equal to that of liquid 11. The enthalpy of fusion of water is 1,435 kcal/mol. The molar entropy change for the melting of ice at 0°C is

 a. 5.26 cal/molx b. 0.526 cal/molK c. 10.52 cal/molx d. 21.04 cal/molt

12. In a reversible isothermal process, the change in internal energy is d. None Dr. Hover RELIEVED !

13. At high pressure Vander Waal's equation reduces to

c. PV-RT-Pb d. PV-RT-a/V b. PV-RT-a/VZ

14. The rate of a chemical reaction doubles for every 10°C rise of temperature. If the temperature is raised by 50°C, the rate of the reaction increases by about

d.64 times c. 32times b. 24 throns a. 10 times

 The rate equation for a reaction: A → B is r = K [A]*. If the initial concentration of the reactant is a mol dm⁻¹, the half life period of the reaction is

d. a/2k c. 2a/k. a. k/a.

16. Half life period of a first order reaction is 1386 seconds. The specific rate constant of the reaction is d. 0.00005 sec⁻¹ c. 0.0005 sec 1 b. 0.005 sec 1 a. 0.05 sec-1

17. SN2 reactions are example of

d. complex reactions b. consecutive reactions c. parallel reactions a. elementary reactions 18. Which of the following would be correct units for the rate constant of a reaction that is second order overall?

c. mol cm⁻³ s⁻¹ d. mol⁻² dm⁴ s⁻¹ b. mol* dm*s*

19. The entropy change of a reaction does not depend on a. the state of reactants and products c. nature of reactants and products

d. initial and final enthalpy change of a reaction. b. different intermediate reactions

20. Acid catalyzed hydrolysis of cane sugar is an example of

b. pseudo 1" order reaction a. zero order reaction

c. 1st order reaction

d. 2nd order reaction.

Roll No: ---

Credit Hours 4(3-1)

d. mean free path

Government College University, Faisalabad External Semester Examination Fall, 2022-23 BS Chemistry (5th Semester) Subjective:

Time: 2 hr 30min Total Marks: (80) CHIM 505

Credit Hours 4(3-1)

Q.2. Define the Maxwells law of molecular velocities. How can we derive and relate different types of molecular velocities in gaseous systems. Also describe the effect of temperature and Molecular wt. on the velocities. (20)

Q 3. Define Free energy and its types, derive a relationship for free energy with equilibrium constant. Also describe the

Q 4. What is the difference between unimolecular and bimolecular reactions. Elaborate with suitable examples. Derive a relationship for Lindmanns theory for unimolecular reactions? Also explain the effect of initial concentration on the order of Q 5. a. How does Virial equation helps to justify the deviation of gases from ideal behavior, which cannot be explained by

b. Define partial molar quantities with suitable examples. What is chemical potential, describe its significance? (10)

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