

11-2 Practice Problems

1. Determine the mass of lithium hydroxide produced when 0.38 g of lithium nitride reacts with water according to the following equation:



$$\frac{0.38\text{g Li}_3\text{N}}{1} \times \frac{1\text{mol Li}_3\text{N}}{35\text{g Li}_3\text{N}} \times \frac{3\text{mol LiOH}}{1\text{mol Li}_3\text{N}} \times \frac{24\text{g LiOH}}{1\text{mol LiOH}} = \boxed{0.78\text{g LiOH}}$$

2. What mass of sodium chloride is produced when chlorine reacts with 0.29 g of sodium iodide?



$$\frac{0.29\text{g NaI}}{1} \times \frac{1\text{mol}}{150} \times \frac{2\text{mol NaCl}}{2\text{mol NaI}} \times \frac{58\text{g}}{1\text{mol}} = \boxed{0.112\text{g NaCl}}$$

3. Determine the mass of carbon dioxide produced when 0.85 g of butane reacts with oxygen according to the following equation:



$$\frac{0.85\text{g C}_4\text{H}_{10}}{1} \times \frac{1\text{mol}}{58\text{g}} \times \frac{8\text{mol CO}_2}{2\text{mol C}_4\text{H}_{10}} \times \frac{44\text{g}}{1\text{mol}} = \boxed{2.58\text{g CO}_2}$$

4. Determine the mass of antimony produced when 0.46 g of antimony(III) oxide reacts with carbon according to the following equation:



$$\frac{0.46\text{g Sb}_2\text{O}_3}{1} \times \frac{1\text{mol}}{291.52\text{g}} \times \frac{2\text{mol Sb}}{1\text{mol Sb}_2\text{O}_3} \times \frac{121.76\text{g}}{1\text{mol}} = \boxed{0.38\text{g Sb}}$$

5. What mass of hydrogen peroxide (H_2O_2) must decompose to produce 0.77 g of water?



$$\frac{0.77\text{g H}_2\text{O}}{1} \times \frac{1\text{mol}}{18\text{g}} \times \frac{2\text{mol H}_2\text{O}_2}{2\text{mol H}_2\text{O}} \times \frac{34\text{g}}{1\text{mol}} = \boxed{1.45\text{g H}_2\text{O}_2}$$

6. What mass of carbon monoxide must react with oxygen to produce 0.69 g of carbon dioxide?



$$\frac{0.69\text{g CO}_2}{1} \times \frac{1\text{mol}}{44\text{g}} \times \frac{2\text{mol CO}}{2\text{mol CO}_2} \times \frac{28\text{g}}{1\text{mol}} = \boxed{0.44\text{g CO}}$$

7. Determine the mass of sodium nitrate produced when 0.73 g of nickel(II) nitrate reacts with sodium hydroxide according to the following equation:



$$\frac{0.73\text{g Ni}(\text{NO}_3)_2}{1} \times \frac{1\text{mol}}{185\text{g}} \times \frac{2\text{mol NaNO}_3}{1\text{mol Ni}(\text{NO}_3)_2} \times \frac{85\text{g}}{1\text{mol}} = \boxed{0.68\text{g NaNO}_3}$$

8. Determine the mass of calcium hydroxide produced when calcium carbide reacts with 0.64 g of water according to the following equation:



$$\frac{0.64\text{g H}_2\text{O}}{1} \times \frac{1\text{mol}}{18\text{g}} \times \frac{1\text{mol Ca}(\text{OH})_2}{2\text{mol H}_2\text{O}} \times \frac{74\text{g}}{1\text{mol}} = \boxed{1.32\text{g Ca}(\text{OH})_2}$$

9. How many grams of ozone (O_3) must decompose to produce 0.87 g of oxygen?



$$\frac{0.87\text{g O}_2}{1} \times \frac{1\text{mol}}{32\text{g O}_2} \times \frac{2\text{mol}}{3\text{mol}} \times \frac{48\text{g O}_3}{1\text{mol}} = \boxed{0.87\text{g O}_3}$$

10. Find the mass of sugar ($\text{C}_6\text{H}_{12}\text{O}_6$) required to produce 1.82 L of carbon dioxide gas at STP from the reaction described by the following equation:



11. How many liters of oxygen are necessary for the combustion of 425 g of sulfur, assuming that the reaction occurs at STP? The balanced equation is $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$.

12. Find the mass of benzene (C_6H_6) required to produce 2.66 L of carbon dioxide gas at STP from the reaction described by the following equation:



14 Prentice Hall Answer Key Chemistry 2014

Stanley E. Manahan



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Environmental Chemistry Stanley E. Manahan, 2022-06-19 With clear explanations real world examples and updated ancillary material the 11th edition of Environmental Chemistry emphasizes the concepts essential to the practice of environmental science technology and chemistry The format and organization popular in preceding editions is used including an approach based upon the five environmental spheres and the relationship of environmental chemistry to the key concepts of sustainability industrial ecology and green chemistry The new edition provides a comprehensive view of key environmental

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Integrated Circuit Fabrication James D. Plummer, Peter B. Griffin, 2023-11-16 Master fundamental technologies for modern semiconductor integrated circuits with this definitive textbook It includes an early introduction of a state of the art CMOS process flow exposes students to big picture thinking from the outset and encourages a practical integration mindset Extensive use of process and TCAD simulation using industry tools such as Silvaco Athena and Victory Process provides students with deeper insight into physical principles and prepares them for applying these tools in a real world setting Accessible framing assumes only a basic background in chemistry physics and mathematics providing a gentle introduction for students from a wide range of backgrounds and over 450 figures many in color and more than 280 end of chapter problems will support and cement student understanding Accompanied by lecture slides and solutions for instructors this is the ideal introduction to semiconductor technology for senior undergraduate and graduate students in electrical engineering materials science and physics and for semiconductor engineering professionals seeking an authoritative introductory reference

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Proceedings of the International Field Exploration and Development Conference 2017 Zhan Qu, Jia'en Lin, 2018-07-11 This book presents selected papers from the 7th International Field Exploration and Development Conference IFEDC 2017 which focus on upstream technologies

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Chemistry for Sustainable Technologies Neil Winterton,2021-02-04 Following the success of the first edition this fully updated and revised book continues to provide an interdisciplinary introduction to

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