Multiple Choice: Entropy

- 1. In which of the following does the entropy decrease?
- A) NaCito Na" Go Ci" Go
- B) 4NO(a) + 6Nb(O(a) + 4NNb(a) + 5Ob(a)
- C) 2NaNOCasis + NasiOCasis + Cosis + NaCisis
- D) CaCOoks + 2HCl (q) + CaClates + COxs) + HaCox

AMS: B

- A) 0.457

B) \$21.40

CO 457

Do 2 70 x 27

may be her

AMS4: C

- 1. Which one of the following reactions has a positive entropy change?
- A) H₀O(g) H₀O₀
- 80 BPost + NHow FaBNHow
- C) 250mm + Onm → 250mm
- DJ May = 3Hays = 2NHays
- ED 22NHaN One 2New + 4HyOre + One

A2005 - 10

1. For a certain process at 1.27°C, $\Delta G = -1.6.20$ kg and $\Delta M = -17.0$ kg. What is the entropy change for this process at this temperature? Express your answer in the form, $\Delta S = -16\%$.

A) 6.3 DK B) +6.3 DK

C) -2.0 J/K

D) +2.0 t/K B) -0.1 t/K

ADMS+ C

Evaluate Δ.S^{*} for the reaction below at 25°C.

CH4(g) + 2CL(g) + CCL(i) + 2H4(g)

	C85.6a3	→ 200 hr (a)	KOYO KARDI	+ 25% date
ARE Alleman	-74.84		48.355.44	4
SOF Brisely	(56.73		14000000	40
A3 -360 5/%	B) -66.9 J/K C) -155 J/K			

A) -360 f/K B) -66.9 f/K D) - 487 f/K B) -387 f/K

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Terry C. Jones

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