

অধ্যায় ১: ল্যাবরেটরির নিরাপদ ব্যবহার



পরীক্ষার কপূর, এখানে সেরে পড়িও, আরও কলেক্সের শীর্ষস্থানীয় কলেক্সমুকের পরীক্ষা এবং মতামত টিউটর পড়িয়ে
ল্যাবর উত্তর অধ্যয়নিক সেএক্স হলেও। এপুণের অনুশীলন করলে তুমি এ অধ্যয়নিক থেকে সেএক্সের সূচনামূলক জ্ঞানমূলক
এক্সের উত্তর লিখতে পারবে সহজেই।

প্রশ্ন ১



পাত্র-A



পাত্র-B

সি. কে. ১০/১ এর ১/১

- মোলের প্রকণ কী?
- "কেলস পানিদ্রুত লবণের প্রাকৃত্য, 'অপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়'— ব্যাখ্যা করো।
- A-পাত্রে বিন্যস্তন এসিডের ভর কত নির্ণয় করো।
- প্রকণ দুটি মিশ্রিত করলে, মিশ্রণটি কী পরিমাণ অক্সিজেন দ্বারা প্রক্রিয় হলে— বিশ্লেষণ করো।

১ নং প্রশ্নের উত্তর

1000 mL প্রকণে কোন প্রকণের 1 mol প্রকৃত প্রকণে উক্ত প্রকণকে মোলের প্রকণ বলে।

কেলস পানিদ্রুত লবণের প্রাকৃত্য অপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়। এর কারণ, আপমাত্রা বৃদ্ধির সাথে সাথে প্রাকৃত্য বৃদ্ধি পেতে থাকলেও একটি নির্দিষ্ট আপমাত্রার কেলস পানি দ্রুত লবণ নিষ্কৃতি হয়। তখন প্রাকৃত্য হ্রাস পেতে থাকে। উদাহরণস্বরূপ $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর কথা বলা যায়। $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর ক্ষেত্রে পানিতে প্রকৃত প্রকণের সময় বিয়োজন প্রক্রিয়াটি অপমাত্রা বৃদ্ধি 32.4°C আপমাত্রার উপরে $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ যখন Na_2SO_4 এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি অপোৎপাদী হয়। এ কারণে 32.4°C আপমাত্রার উপরে আপমাত্রা বৃদ্ধিজনিত কারণে প্রাকৃত্য হ্রাস ঘটে।

২

আমরা জানি,
 $w = SMV$
 $= (0.5 \times 98 \times 0.25) \text{ g}$
 $= 12.25 \text{ g}$

অতএব, প্রকৃত H_2SO_4 এর ভর
 $= 12.25 \text{ g}$

উদীপকের A-পাত্রে অবস্থিত,
 H_2SO_4 এর আণবিক ভর,
 $M = 1 \times 2 + 32 + 16 \times 4$
 $= 2 + 32 + 64$
 $= 98 \text{ g/mol}$

প্রকণের ঘনমাত্রা,
 $S = 0.5 \text{ M}$
প্রকণের আয়তন,
 $V = 250 \text{ mL}$
 $= \frac{250}{1000} = 0.25 \text{ L}$

প্রকৃত H_2SO_4 এর ভর, $w = ?$

প্রকণ-A: H_2SO_4 এর মোল সংখ্যা $= \left(0.5 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.125 \text{ mol}$

প্রকণ-B: Na_2CO_3 এর মোল সংখ্যা $= \left(0.11 \times \frac{700}{1000}\right) \text{ mol}$
 $= 0.077 \text{ mol}$

সম্প্রতিক বিক্রিয়া: $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$

সমীকরণ হতে পাই, 1 mol $\text{H}_2\text{SO}_4 = 1 \text{ mol Na}_2\text{CO}_3$

1 mol Na_2CO_3 বিক্রিয়া করে H_2SO_4 এর 1 mol এর সাথে
 $0.077 \text{ mol} = 0.077 \text{ mol} = 0.077 \text{ mol} = 0.077 \text{ mol}$

অবশিষ্ট H_2SO_4 এর পরিমাণ $= (0.125 - 0.077) \text{ mol}$
 $= 0.048 \text{ mol}$

পাত্র-A এর প্রকণের সাথে পাত্র-B এর প্রকণের বিক্রিয়ার ফলে 0.048 mol H_2SO_4 অবশিষ্ট থাকায় মিশ্রিত প্রকণের প্রকৃতি হবে অক্সিজেন।

প্রশ্ন ২ এককাল ছাত্র ল্যাবরেটরিতে X-পদ্ধতিতে 10 mg এর প্রকণ প্রকৃতি এবং Y-পদ্ধতিতে 0.5 g পরিমাণ 2% $\frac{w}{v}$ HCl এর প্রকণ প্রকৃতি কত।

সি. কে. ১০/১ এর ১/১

- নোড কাকে বলে?
- Cl_2 অপোলায়— ব্যাখ্যা করো।
- উদীপকের এসিডটির মোলার ঘনমাত্রা নির্ণয় করো।
- উদীপকের কোন পদ্ধতি পরিবেশবান্ধব? বিশ্লেষণ করো।

২ নং প্রশ্নের উত্তর

নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেক্ট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে নোড বলে।

সমযোজী বন্ধনে আবদ্ধ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা সমান হলে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হলে সমযোজী বন্ধন গঠিত হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর কম হলে অপোলায় হয়। যেহেতু Cl এর তড়িৎ ঋণাত্মকতা 3.0 ও C অপুতে তড়িৎ ঋণাত্মকতার পার্থক্য $(3 - 3) = 0$, তাই Cl_2 বিশুদ্ধ সমযোজী ও অপোলায় যৌগ। উল্লেখ্য যৌগিক অণুর ক্ষেত্রে এধরন থাকে।

উদীপকে বর্ণিত এসিডটি হলো 2% $\left(\frac{w}{v}\right) \text{ HCl}$

2% $\left(\frac{w}{v}\right) \text{ HCl}$ এর অর্থ হলো—

100 mL প্রকণে HCl এর ভর 2g

∴ 1000 mL " " " " $\left(\frac{2 \times 1000}{100}\right) \text{ g}$
 $= 20 \text{ g}$

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Arshad Iqbal



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